



# Compaq Evo N1desktop Series Manual

**RD Boyd**



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web jan 29 2020 the ionic form of this reaction is 2 na aq 2 cl aq cu 2 aq so 4<sup>2-</sup> aq 2 na aq so 4<sup>2-</sup> aq cucl 2 s the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six no 3 aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation 3ba 2 aq 2po 4 3 aq rightarrow ba 3 po 4 2 s

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changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a **practice problems on net ionic equations chem kmacgill com** - Sep 12 2023

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 $\text{AgNO}_3 \text{ aq} + \text{KCl aq} \rightarrow \text{AgCl s} + \text{KNO}_3 \text{ aq}$  total ionic  $\text{Ag}^+ \text{ aq} + \text{NO}_3^- \text{ aq} + \text{K}^+ \text{ aq} + \text{Cl}^- \text{ aq} \rightarrow \text{AgCl s} + \text{K}^+ \text{ aq} + \text{NO}_3^- \text{ aq}$  net ionic  $\text{Ag}^+ \text{ aq} + \text{Cl}^- \text{ aq} \rightarrow \text{AgCl s}$  2

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+ \text{ aq} + \text{OH}^- \text{ aq} \rightarrow \text{H}_2\text{O l}$  the spectator ions in  $\text{KOH aq} + \text{HNO}_3 \text{ aq} \rightarrow \text{KNO}_3 \text{ aq} + \text{H}_2\text{O l}$  are  $\text{K}^+ \text{ aq}$  and  $\text{NO}_3^- \text{ aq}$  how to identify and list the

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web the  $\text{Ag}^+ \text{ aq}$  and  $\text{Cl}^- \text{ aq}$  ions become  $\text{AgCl s}$  but the  $\text{Na}^+ \text{ aq}$  ions and the  $\text{NO}_3^- \text{ aq}$  ions stay as  $\text{Na}^+ \text{ aq}$  ions and  $\text{NO}_3^- \text{ aq}$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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