

GCSE Chemistry: Covalent Bonding Lesson & Activities

Covalent Bonding

Specification Link:
Chemical bonding: 4.1.1.4

Task

Draw a diagram to show the electrons and the covalent bonds in the molecules below:

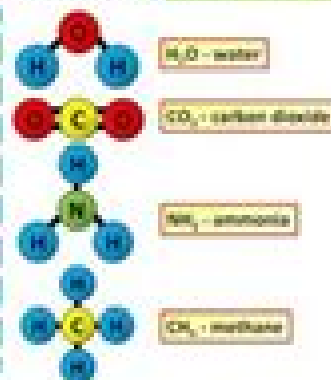
- | | |
|-----------------------------------|--|
| 1. CO_2 - carbon dioxide | 7. PCl_5 - phosphorus pentachloride |
| 2. CH_4 - methane | 8. HCl - hydrogen chloride |
| 3. NH_3 - ammonia | 9. SiF_4 - silicon tetrafluoride |
| 4. H_2O - water | 10. NO_2 - nitrogen dioxide |
| 5. O_2 - oxygen | 11. SiO_2 - silicon dioxide |
| 6. H_2 - hydrogen | 12. CBr_4 - carbon tetrabromide |

Simple Covalent Compounds

Specification Link:
Chemical bonding: 4.1.1.4

Introduction

Hydrogen, ammonia, methane and water are all simple molecules with covalent bonds. All have very strong bonds between the atoms, but much weaker forces holding the molecules together. When one of these substances melts or boils, it is these weak 'intermolecular forces' that break, not the strong covalent bonds.



Giant Covalent Compounds

Specification Link:
Chemical bonding: 4.1.1.4

Introduction

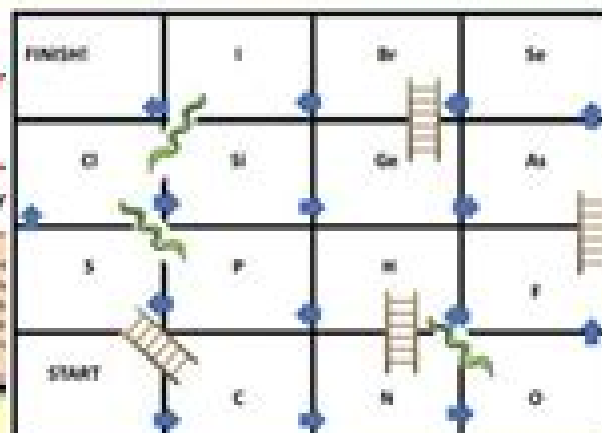
Giant covalent structures contain a lot of non-metal atoms, each joined to adjacent atoms by covalent bonds. The atoms are usually arranged into giant regular lattices - extremely strong structures because of the many bonds involved.

Properties of giant covalent structures

- Very high melting points - Substances with giant covalent structures have very high melting points, because a lot of strong covalent bonds must be broken.
- Variable conductivity - Diamond does not conduct electricity. Graphite contains free electrons, so it does conduct electricity.

Task

1. Read the giant covalent structures information slide.
2. In pairs, plan a role-play on giant covalent structures.
3. Be ready to pitch to others.



Simple molecular compounds SPAG activity

Covalent bonds form between non-metal atoms. Each bond consists of a shared pair of electrons and is very strong. Covalently bonded substances fall into two main types:

- Simple molecules
- Giant covalent structures

Properties of simple molecular substances

Low melting and boiling points - This is because the weak intermolecular forces break down easily.

Non-conductive - Substances with a simple molecular structure do not conduct electricity. This is because they do not have any free electrons or an overall electric charge.

Hydrogen, ammonia, methane and water are all simple molecules with covalent bonds. The strong covalent bonds between the atoms hold the molecules together. When one of these substances melts or boils, it is these weak 'intermolecular forces' that break, not the strong covalent bonds. Simple molecular substances are usually liquids or solids with low melting and boiling points.

Science

Covalent Bonds Lesson

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Covalent Bonds Lesson :

BIOLOGY | LIFE | Unit 1 | From Atoms to Cells Smart Biology Inc.,2022-05-01 BIOLOGY LIFE UNIT 1 From Atoms to Cells focuses on the physical components that make up cells Embark on one continuous journey to understand and appreciate the interconnections between the subatomic atomic molecular macromolecular and cellular worlds We spend time covering the basics so you can understand the complex Moreover we explain the underlying why questions so you can truly understand This downloadable e book includes access to text over 350 high quality accurate figures 40 interactive structures and more to suit all learners Note Access to our online courseware including our animated video lessons is not included in this eBook but can be purchased at [www smart biology com](http://www.smartbiology.com) E3 Chemistry Guided Study Book - 2018 Home Edition (Answer Key Included) Effiong Eyo,2017-12-08 Chemistry students and Homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Guided Study Book 2018 With E3 Chemistry Guided Study Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with guided step by step solutions to study and follow Practice multiple choice and short answer questions along side each concept to immediately test student understanding of the concept 12 topics of Regents question sets and 2 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN 978 1979088374 The Home Edition contains answer key to all questions in the book Teachers who want to recommend our Guided Study Book to their students should recommend the Home Edition Students and and parents whose school is not using the Guided Study Book as instructional material as well as homeschoolers should also buy the Home edition The School Edition does not have the answer key in the book A separate answer key booklet is provided to teachers with a class order of the book Whether you are using the school or Home Edition our E3 Chemistry Guided Study Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year PLEASE NOTE Although reading contents in both the school and home editions are identical there are slight differences in question numbers choices and pages between the two editions Students whose school is using the Guided Study Book as instructional material SHOULD NOT buy the Home Edition Also available in paperback print *E-chemistry Iii Tm (science and Technology)' 2003 Ed. ,* **A Teaching Guide to Revitalizing STEM Education** Daryao Khatri,2012-11-13 Standard math and science textbooks typically follow a deductive style of content presentation that involves too much lecturing too much of the teacher s back at the chalkboard too little interaction with students and too little time for all of the students to take adequate notes By reading and using A Teaching Guide to Revitalizing STEM Education educators will rediscover how to streamline the subject matter math physics statistics and organic chemistry by eliminating unnecessary difficulties and

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