

Name: _____

Sci # _____

COVALENT BOND WEBQUEST

DIRECTIONS: Answer the following by going to the notes website. You can go to Mrs. Slowik's website and use the links or type in the addresses for each section if needed.

Activity #1- Introduction to Covalent Bonding

Open [Chemical Bonding](#). Scroll down to the heading "Covalent Bonding." OR
http://www.visionlearning.com/library/module_viewer.php?mid=55

1. As opposed to _____ bonding in which a complete transfer of electrons occur, _____ bonding occurs when two or more elements _____ electrons. Covalent bonding occurs because the atoms in the compound have a similar tendency for electrons (generally to _____ electrons). This most commonly occurs when two _____ bond together. Because both of the nonmetals will want to _____ electrons, the elements involved will share electrons in an effort to _____ their valence shells.

2. How many valence electrons are in one atom of hydrogen?

3. How many valence electrons does hydrogen need to have a full first shell?

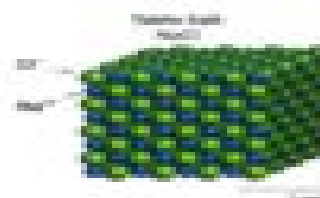
4. How does the hydrogen atom "pick up" another electron?

5. What compound does hydrogen form?

6. How do hydrogen atoms make a covalent bond?

7. Visit the simulation, [Covalent bonding between hydrogen atoms](#). OR
http://www.visionlearning.com/library/flash_viewer.php?fid=1346&mid=55
Describe or draw what you see.

8. Label the substances below as "ionic" or "covalent".



Covalent Bond Webquest Key

Daniel Berg



Covalent Bond Webquest Key:

Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more **Chemistry of the Covalent Bond** Leallyn Burr Clapp, 1957 *The Covalent Bond* Henry Sinclair Pickering, 1977 **Chemistry of the Covalent Bond** Leallyn B. Clapp, 1950 The Binding Force Daniel Berg, 1966 The Covalent Bond M. A. Kazi, 1983 *Chemical Bonding* Source Wikipedia, 2013-09 Please note that the content of this book primarily consists of articles available from Wikipedia or other free sources online Pages 157 Chapters Covalent bond Hydrogen bond Chemical bond Intermolecular force Electronegativity Ionic bond Pauli exclusion principle Molecular orbital Metallic bond Disulfide bond Atomic orbital Dipolar bond Ligand Peptide bond Van der Waals force Mesomeric effect Electron counting Cooperative binding Covalent radius Aromaticity Hypervalent molecule Halogen bond Quantum mechanical explanation of intermolecular interactions Molecular orbital diagram Radical Hyperconjugation Resonance Polyhedral skeletal electron pair theory Orbital hybridisation Electron affinity Crystal field theory Anodic bonding Ruggero Santilli Salt bridge Non innocent ligand Chemical polarity Lewis structure Valence Valence electron Quintuple bond Coordination geometry Bonding in solids Walsh diagram Hapticity 18 Electron rule Electrovalency Ligand field theory Conjugated system Lennard Jones potential Bent s rule Inert pair effect Atoms in molecules Cation pi interaction Jahn Teller effect Dihydrogen complex Valence bond theory Bent bond Morse

potential Hydrophobic effect Stacking Lanthanide contraction Formal charge Octet rule Entropic force London dispersion
 force Slater's rules Coordination number Metal ligand multiple bond Auophilicity Linear combination of atomic orbitals
 molecular orbital method Agostic interaction Inductive effect Electron localization function Intercalation Double bond rule
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 Phillips, CHICAGO UNIV ILL., 1967 The general theory of covalent bonding developed in AD 657 432 is extended to
 heteropolar crystals In diatomic crystals the theory depends on two parameters an energy gap for each atom or an energy
 gap for the medium and a polarizability parameter η The latter can be determined by comparison with empirical
 pseudopotential form factors the value obtained for GaAs and ZnSe which brings the form factors of the latter into agreement
 with the ion core form factors of Animalu and Heine is $\eta = 0.05$ or 0.05 This polarizability parameter can be used to
 estimate the effective charge e_c defined by Callen which describes the splitting of longitudinal and transverse optic modes of
 vibration at $q = 0$ The result in GaAs is $e_c = 0.14$ or 0.04 in rough agreement with the experimental value of 0.20 or 0.03
 The experimental values for e_c may be used to estimate the gap parameters for each atom when complete data are
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