

# Hurricane Preparation Checklist



## Water

at least 1 gallon daily per person for 3 to 7 days



## Food

at least enough for 3 to 7 days



- Non-perishable packaged or canned food/ juices
- Foods for infants or the elderly
- Snack foods
- Non-electric can opener
- Cooking tools/fuel
- Paper plates/plastic utensils



## Blankets Pillows



## Clothing

seasonal/rain gear/ sturdy shoes



## First Aid Kit

medicines  
prescription drugs



## Special Items

for babies and the elderly



## Flashlight Batteries



## Telephones

Fully charged cell phone with extra battery and a traditional telephone set



## Toiletries

hygiene items  
moisture wipes



## Radio

Battery operated and NOAA weather radio

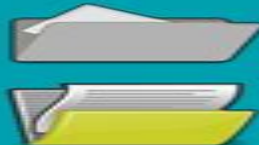


## Cash & Credit Cards

Banks and ATMs may not be available for extended periods



## Keys



## Important Documents

in a waterproof container - insurance, medical records, bank account numbers, social security cards



## Toys, Books, Games



## Tools

keep a set with you during the storm



## Pet Care Items

- proper identification/ medications
- ample supply of food and water
- a carrier or cage
- muzzle or leash



## Vehicle fuel tanks filled

# Data Center Hurricane Preparation Guide

**Pat J. Fitzpatrick**



## **Data Center Hurricane Preparation Guide:**

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 Presents background information on issues people organizations and publications related to hurricanes and provides guidance on where additional information may be found about a specific topic

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explanation the entire reaction was multiplied by 2 therefore  $\Delta H$  must also be halved 3 answer c explanation option c is  
incorrect because 566 kj mol is the  $\Delta H$  per mole of  $O_2$  and not for 2 moles

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an endothermic process it requires 6.01 kJ to melt one mole of ice  $H_2O(s)$  the enthalpy value will change if the number of  
moles varies from the 1 1 reaction stoichiometry this is an exothermic process



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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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