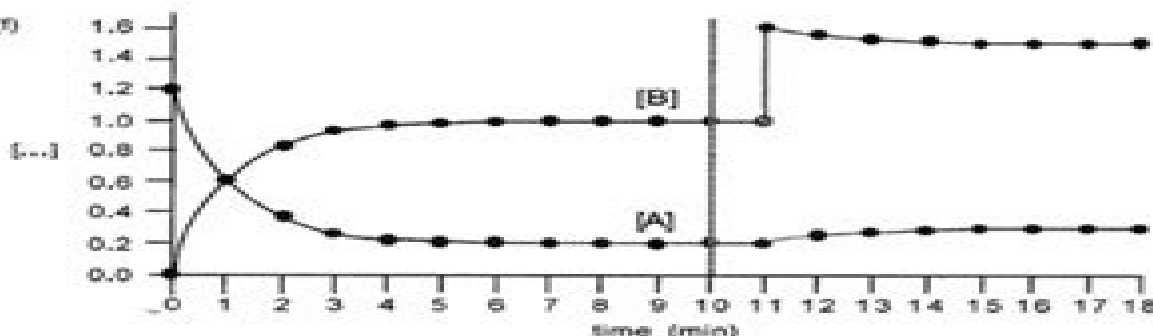


## ANSWERS TO UNIT II : EQUILIBRIUM

1. (a) Wrap the container in insulation.  
(b) Wrap the container in aluminum foil or paint it with black paint.  
(c) Put the system in an airtight container.
2. (a) The colour does not change, so  $\text{NO}_2(\text{g})$  is being made at the same rate that it is destroyed.  
(b) Temperature CAN affect an equilibrium – the colour became lighter or darker when the temperature was changed, meaning more or less  $\text{NO}_2(\text{g})$  was present.  
(c) The colour does not change while the tube full of gas remains at a constant  $100^\circ\text{C}$ . The colour would become very dark red–brown if the temperature were raised above  $100^\circ\text{C}$ .  
(d) The reaction is endothermic as written:  $\text{N}_2\text{O}_4(\text{g}) + \text{energy} \rightleftharpoons 2 \text{NO}_2(\text{g})$ . As heat is added the forward reaction should occur to a greater extent and produce more of the red–brown  $\text{NO}_2(\text{g})$ , which is exactly what occurred.  
(e)  $\text{N}_2\text{O}_4(\text{g})$  predominated at low temperatures (colourless).  
 $\text{NO}_2(\text{g})$  predominated at high temperatures (dark red–brown).  
At room temperature the content of the tube was a mixture of  $\text{N}_2\text{O}_4(\text{g})$  and  $\text{NO}_2(\text{g})$ .  
(f) The tubes should become the same colour. A tube containing mostly  $\text{N}_2\text{O}_4(\text{g})$  at low temperatures and another tube containing mostly  $\text{NO}_2(\text{g})$  at high temperatures eventually became the same colour at room temperature.
3. The system is NOT at equilibrium because the system is OPEN. Gas is steadily escaping from the liquid but no gas is going back into the liquid.
4. Water is evaporating inside the flask, but water vapour is re–entering the liquid phase at the same rate so that the system is at equilibrium.
5. The reaction will not make pure  $\text{COCl}_2(\text{g})$  because there will always be some  $\text{CO}(\text{g})$  and  $\text{Cl}_2(\text{g})$  in equilibrium with the  $\text{COCl}_2(\text{g})$ . If some pure  $\text{COCl}_2(\text{g})$  could be obtained, it would partly decompose into  $\text{CO}(\text{g})$  and  $\text{Cl}_2(\text{g})$  as equilibrium occurred.

6. (a), and 7(f)



- (b) Equilibrium occurs by 9 min. At equilibrium the concentrations of both A and B remain constant. The forward and reverse rates remain constant and the forward rate equals the reverse rate.
- (c) After 1 minute,  $[\text{A}] = [\text{B}]$ ; that is,  $[\text{REACTANT}] = [\text{PRODUCT}]$ . However, even though the reaction equation shows that one molecule of A produces one molecule of B, it is not required (and it is not true) that  $[\text{REACTANT}] = [\text{PRODUCT}]$  at equilibrium (although it is remotely possible).
- (d) Forward rate is greatest when  $[\text{A}]$  is greatest. As  $[\text{A}]$  decreases, so does forward rate.
- (e)  $\text{Ratio} = \frac{1.000}{0.200} = 5.00$

# Data Sheet Equil039introducing Equilibrium With Answers

**E Durkheim**



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## **Table of Contents Data Sheet Equ039introducing Equilibrium With Answers**

1. Understanding the eBook Data Sheet Equ039introducing Equilibrium With Answers
  - The Rise of Digital Reading Data Sheet Equ039introducing Equilibrium With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Data Sheet Equ039introducing Equilibrium With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Data Sheet Equ039introducing Equilibrium With Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Data Sheet Equ039introducing Equilibrium With Answers
  - Personalized Recommendations
  - Data Sheet Equ039introducing Equilibrium With Answers User Reviews and Ratings
  - Data Sheet Equ039introducing Equilibrium With Answers and Bestseller Lists
5. Accessing Data Sheet Equ039introducing Equilibrium With Answers Free and Paid eBooks
  - Data Sheet Equ039introducing Equilibrium With Answers Public Domain eBooks
  - Data Sheet Equ039introducing Equilibrium With Answers eBook Subscription Services

- Data Sheet Equ039introducing Equilibrium With Answers Budget-Friendly Options
- 6. Navigating Data Sheet Equ039introducing Equilibrium With Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Data Sheet Equ039introducing Equilibrium With Answers Compatibility with Devices
  - Data Sheet Equ039introducing Equilibrium With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Data Sheet Equ039introducing Equilibrium With Answers
  - Highlighting and Note-Taking Data Sheet Equ039introducing Equilibrium With Answers
  - Interactive Elements Data Sheet Equ039introducing Equilibrium With Answers
- 8. Staying Engaged with Data Sheet Equ039introducing Equilibrium With Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Data Sheet Equ039introducing Equilibrium With Answers
- 9. Balancing eBooks and Physical Books Data Sheet Equ039introducing Equilibrium With Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Data Sheet Equ039introducing Equilibrium With Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Data Sheet Equ039introducing Equilibrium With Answers
  - Setting Reading Goals Data Sheet Equ039introducing Equilibrium With Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Data Sheet Equ039introducing Equilibrium With Answers
  - Fact-Checking eBook Content of Data Sheet Equ039introducing Equilibrium With Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
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