

Class 11 Chemical Bonding and Molecular Structure Important Questions with Answers

Short Answer Type Questions

1. Explain the non linear shape of H_2S and non planar shape of PCl_3 using valence shell electron pair repulsion theory.

Answer.

- The main atom in H_2S is S, which has two lone pairs. These lone pairs cause repulsion and displace the H-S bond, resulting in a non-linear shape.



- PCl_3 has a trigonal planar structure. P has three single bonds and one lone pair (pair of unshared electrons). Each chlorine atom has a single 3p orbital that is completely occupied. The overlap of a phosphorus sp^3 hybrid orbital with a singly occupied chlorine 3p orbital results in the formation of P-Cl bonds. Three lone pairs are held by each Cl atom.



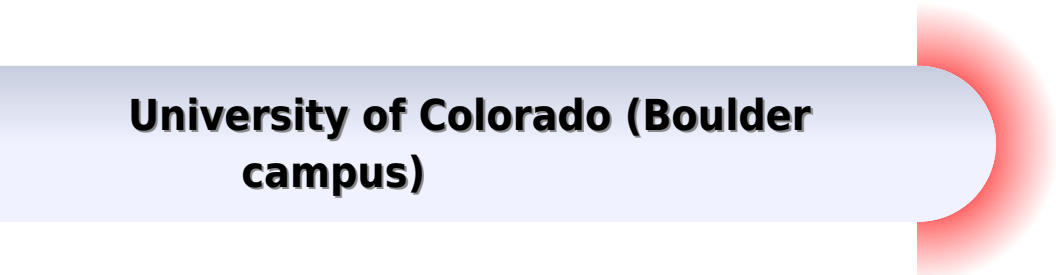
2. Using molecular orbital theory, compare the bond energy and magnetic character of O_2^+ and O_2^- species.

Answer.

The electronic configurations of O_2^+ and O_2^- species are as follows:

Covalent Bonding And Molecular Structure Lab Answers

**University of Colorado (Boulder
campus)**



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United States Air Force Academy United States Air Force Academy, 1968 **Annual Catalog - United States Air Force Academy** United States Air Force Academy, 1971

Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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