

Parts of the Atom

**Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit.)*

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS
proton	proton	Nucleus	2000 X More than an e ⁻
neutron	neutron	Nucleus	2000 X More than an e ⁻
electron	electron	Orbitals / Energy Levels	1

**Here's how to predict the structure of any element on the periodic table. For example, look at carbon:*

6
C
carbon
12.011

Error: Reference source not foundatomic number: This is the number of protons (also the number of electrons)

Error: Reference source not foundchemical symbol: This is an abbreviation for the element.

Error: Reference source not foundelement name: This is the name of the element.

Error: Reference source not foundatomic mass: This number (rounded off) tells you the number of protons *plus* the number of neutrons. (Electrons are too small to be included.)

To find the number of protons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of protons: 6

To find the number of neutrons, subtract the atomic number from the atomic mass:

Example for carbon:

Atomic mass: 12 (rounded)

Number of neutrons: $12 - 6 = 6$

To find the number of electrons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of electrons: 6

**Complete the table for the elements with atomic numbers 1-10.*

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

Cstephenmurray The Atom Answers

Harold Albert Wilson



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The Structure of the Atom Stephen Miall,1922 *The Atom Uncovered* Mark Andrew Newstead,Stephen Charles Newstead,2010 In this book we show how electromagnetic waves can form the sub atomic particles i e the proton neutron and electron as well as how these can fit together to form the nuclei of atoms and of all their isotopes From this we are able to combine the weak and strong force into the electromagnetic force *The Question of the Atom* Mary J. Nye,1983-01-01

The Structure of the Atom Stephen Miall,1922 **The Britannica Guide to the Atom** Britannica Educational Publishing,2010-10-01 Common to all matter and its smallest and most fundamental unit the atom determines the properties of a substance and how it responds to other materials and stimuli Dividing the atom further reveals several smaller particles whose activity forms the heart of nuclear physics Readers will discover the structure of the atom as well as the ways in which its electrons protons and neutrons facilitate both radioactivity and nuclear reactions both of which have been crucial to technological advancements since the 20th century Explaining the Atom Selig Hecht,1947 This is a book for the complete layman It follows the steps that were taken historically in going from the earliest questions about the nature of common substances to the large scale liberation of atomic energy The number of these steps is small and the ideas involved are simple They can be described in all their essentials without assuming any knowledge of physics chemistry and mathematics Preface

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