

Stoichiometry

Section 11.1 Defining Stoichiometry

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Practice Problems

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1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule N_2 + 3 molecules $\text{H}_2 \rightarrow$ 2 molecules NH_3

1 mole N_2 + 3 moles $\text{H}_2 \rightarrow$ 2 moles NH_3

Mass:

N_2 : (2 mol)(14.007 g/mol) = 28.014 g

3H_2 : (6 mol)(1.008 g/mol) = 6.048 g

2NH_3 : (2 mol)(14.007 g/mol) +
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g N_2 + 6.048 g $\text{H}_2 \rightarrow$ 34.062 g NH_3

34.062 g reactants = 34.062 g products



1 molecule HCl + 1 formula unit $\text{KOH} \rightarrow$
1 formula unit KCl + 1 molecule H_2O

1 mole HCl + 1 mole $\text{KOH} \rightarrow$
1 mole KCl + 1 mole H_2O

Mass:

HCl : (1 mol)(1.008 g/mol) +
(1 mol)(35.453 g/mol) = 36.461 g

KOH : (1 mol)(39.098 g/mol) +
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)
= 56.105 g

KCl : (1 mol)(39.098 g/mol) +
(1 mol)(35.453 g/mol) = 74.551 g

H_2O : (2 mol)(1.008 g/mol) +
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g HCl + 56.105 g $\text{KOH} \rightarrow$
74.551 g KCl + 18.015 g H_2O

92.566 g reactants = 92.566 g products



2 atoms Mg + 1 molecule $\text{O}_2 \rightarrow$ 2 formula units MgO

2 moles Mg + 1 mole $\text{O}_2 \rightarrow$ 2 moles MgO

Mass:

2Mg : (2 mol)(24.305 g/mol) = 48.610 g

O_2 : (2 mol)(15.999 g/mol) = 31.998 g

2MgO : (2 mol)(24.305 g/mol) +
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g Mg + 31.998 g $\text{O}_2 \rightarrow$ 80.608 g MgO

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



$2\text{Na}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$

2 atoms Na + 2 molecules $\text{H}_2\text{O} \rightarrow$
2 formula units NaOH + 1 molecule H_2

2 mol Na + 2 mol $\text{H}_2\text{O} \rightarrow$
2 mol NaOH + 1 mol H_2

Mass:

2Na : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$: (4 mol)(1.008 g/mol) +
(2 mol)(15.999 g/mol) = 36.030 g

2NaOH : (2 mol)(22.990 g/mol) +
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)
= 79.994 g

H_2 : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g Na + 36.030 g $\text{H}_2\text{O} \rightarrow$
79.994 g NaOH + 2.016 g H_2

82.01 g reactants = 82.01 g products

Chapter 11 Stoichiometry Study Guide Answers

Julie R. Frentrup



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