Stoichiometry

Section 11.1 Defining Stoichiometry

pages 368-372

Practice Problems

pages 371-372

- Interpret the following balanced chemical equations in terms of particles, moles, and mass.
 Show that the law of conservation of mass is observed.
 - a. $N_s(g) + 3H_s(g) \rightarrow 2NH_s(g)$

1 molecule N₂ + 3 molecules H₂ → 2 molecules NH₃

1 mole N₂ + 3 moles H₂ -> 2 moles NH₃

Milance

N₂: (2 mol)(14.007 g/mol) = 28.014 g

3H₂: (6 mol)(1.008 g/mol) = 6.048 g

2NH₃: (2 mol)(14.007 g/mol) + (6 mol)(1.008 g/mol) = 34.062 g

28.014 g N₂ + 6.048 g H₂ -> 34.062 g NH₃

34.062 g reactants = 34.062 g products

b. HCl(aq) + KOH(aq) → KCl(aq) + H₂O(I)

1 molecule HCl + 1 formula unit KOH → 1 formula unit KCl + 1 molecule H₂O

1 mole HCI + 1 mole KOH →

1 mole KCl + 1 mole H₂O

Politica execu-

HCl: (1 mol)(1.008 g/mol) + (1 mol)(35.453 g/mol) = 36.461 g

KOH: (1 mol)(39.098 g/mol) + (1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol) = 56.105 g

KCl: (1 mol)(39.098 g/mol) + (1 mol)(35.453 g/mol) = 74.551 g

 H_2O : (2 mol)(1.008 g/mol) + (1 mol)(15.999 g/mol) = 18.015 g

36.461 g HCl + 56.105 g KOH --74.551 g KCl + 18.015 g H₂O

92.566 g reactants = 92.566 g products

c. $2Mg(s) + O_s(g) \rightarrow 2MgO(s)$

2 atoms Mg + 1 molecule O₁ → 2 formula units MgO

2 moles Mg + 1 mole O₂ → 2 moles MgO

Marks

2Mg: (2 mol)(24.305 g/mol) = 48.610 g

O₂: (2 mol)(15.999 g/mol) = 31.998 g

2MgO: (2 mol)(24.305 g/mol) + (2 mol)(15.999 g/mol) = 80.608 g

48.610 g Mg + 31.998 g O₂ → 80.608 g MgO

80.608 g reactants = 80.608 g products

- Challenge For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.
 - a. __Na(s) + __H₂O(l) → __NaOH(aq) + __H₂(g)

 $2Na(s) + 2H_sO(l) \rightarrow 2NaOH(aq) + 1H_s(q)$

2 atoms Na + 2 molecules H₂O --2 formula units NaOH + 1 molecule H₄

2 mol Na + 2 mol H₂O → 2 mol NaOH + 1 mol H₂

Maga:

2Na: (2 mol)(22.990 g/mol) = 45.980 g

2H₂O: (4 mol)(1.008 g/mol) + (2 mol)(15.999 g/mol) = 36.030 g

2NaOH: (2 mol)(22.990 g/mol) + (2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol) = 79.994 g

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 H_s : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g Na + 36.030 g H₂O → 79.994 g NaOH + 2.016 g H₂

82.01 g reactants = 82.01 g products

Chapter 11 Study Guide Chemistry Stoichiometry Answer Key

Steven S. Zumdahl, Paul B. Kelter

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the test Familiarizing yourself with the directions and format of the exam will not only save you time but will also ensure that you are familiar enough with the SAT II Chemistry Subject Test to avoid nervousness and the mistakes caused by being nervous Do your scratchwork in the margins of the test booklet You will not be given scrap paper during the exam and you may not perform scratchwork on your answer sheet Space is provided in your test booklet to do any necessary work or draw diagrams If you are unsure of an answer guess However if you do guess guess wisely Use the process of elimination by going through each answer to a question and ruling out as many of the answer choices as possible By eliminating three answer choices you give yourself a fifty fifty chance of answering correctly since there will only be two choices left from which to make your guess Mark your answers in the appropriate spaces on the answer sheet Fill in the oval that corresponds to your answer darkly completely and neatly You can change your answer but remember to completely erase your old answer Any stray lines or unnecessary marks may cause the machine to score your answer incorrectly When you have finished working on a section you may want to go back and check to make sure your answers correspond to the correct questions Marking one answer in the wrong space will throw off the rest of your test whether it is graded by machine or by hand You don't have to answer every question You are not penalized if you do not answer every question. The only penalty results from answering a question incorrectly Try to use the guessing strategy but if you are truly stumped by a question remember that you do not have to answer it Work quickly and steadily You have a limited amount of time to work on each section so you need to work quickly and steadily Avoid focusing on one problem for too long Before the Test Make sure you know where your test center is well in advance of your test day so you do not get lost on the day of the test On the night before the test gather together the materials you will need the next day Your admission ticket Two forms of identification e g driver s license student identification card or current alien registration card Two No 2 pencils with erasers Directions to the test center A watch if you wish but not one that makes noise as it may disturb other test takers On the day of the test you should wake up early after a good night's rest and have breakfast Dress comfortably so that you are not distracted by being too hot or too cold while taking the test Also plan to arrive at the test center early This will allow you to collect your thoughts and relax before the test and will also spare you the stress of being late If you arrive after the test begins you will not be admitted to the test center and you will not receive a refund During the Test When you arrive at the test center try to find a seat where you feel most comfortable Follow all the rules and instructions given by the test supervisor If you do not you risk being dismissed from the test and having your scores canceled Once all the test materials are passed out the test instructor will give you directions for filling out your answer sheet Fill this sheet out carefully since this information will appear on your score report After the Test When you have completed the SAT II Chemistry Subject Test you may hand in your test materials and leave Then go home and relax When Will I Receive My Score Report and What Will It Look Like You should receive your score report about five weeks after you take the test This report will include your scores percentile ranks and interpretive

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