

Mixtures and Solutions

Section 14.1 Types of Mixtures

pages 476–479

Section Assessment 14.1

page 479

- 1. Explain** Use the properties of seawater to describe the characteristics of mixtures.

Answers will vary but might include that seawater is a heterogeneous mixture with dirt and mud particles, and it is a homogeneous mixture with dissolved substances.

- 2. Distinguish** between suspensions and colloids.

Suspension particles are larger than colloidal particles. Suspension particles settle out of the mixture, whereas colloidal particles do not.

- 3. Identify** the various types of solutions. Describe the characteristics of each type of solution.

All solutions are homogeneous mixtures containing two or more substances. Solutions may be liquid, solid, or gas. Solution types are identified in Table 14.2.

- 4. Explain** Use the Tyndall effect to explain why it is more difficult to drive through fog using high beams than using low beams.

High beams are aimed farther down the road than low beams. Because the fog scatters light, there is less light from the high beams to illuminate the road than from the low beams. Also, because the high beams are aimed more directly into the fog, more of their light is reflected back toward the driver, making it more difficult to see.

- 5. Describe** different types of colloids.

See Table 14.1 for descriptions of colloid types.

- 6. Explain** Why do dispersed colloid particles stay dispersed?

The particles do not settle out because they have polar or charged layers surrounding them. These layers repel each other and prevent the particles from settling or separating.

- 7. Summarize** What causes Brownian motion?

Collisions of particles of the dispersion medium with the dispersed particles results in Brownian motion.

- 8. Compare and Contrast** Make a table that compares the properties of solutions, suspensions, and colloids.

Student tables will vary, but should include particle size, if the particles settle out, and if the particles display the Tyndall effect. A sample table follows.

Suspensions, Colloids, and Solutions			
	Particle size	Particles settle?	Tyndall effect?
Suspensions	Large (wide variation)	Yes	Yes
Colloids	1 nm–1000 nm	No	Yes
Solutions	Atomic scale (atoms, ions, and molecules)	No	No

Section 14.2 Solution Concentration

pages 480–488

Practice Problems

pages 481–488

- 9.** What is the percent by mass of NaHCO_3 in a solution containing 20.0 g NaHCO_3 dissolved in 600.0 mL H_2O ?

$$600.0 \text{ mL } \text{H}_2\text{O} \times 1.0 \text{ g/mL} = 600.0 \text{ g } \text{H}_2\text{O}$$

$$\frac{20 \text{ g } \text{NaHCO}_3}{600 \text{ g } \text{H}_2\text{O} + 20 \text{ g } \text{NaHCO}_3} \times 100 = 3\%$$

- 10.** You have 1500.0 g of a bleach solution. The percent by mass of the solute sodium hypochlorite, NaOCl , is 3.62%. How many grams of NaOCl are in the solution?

$$3.62\% = 100 \times \frac{\text{mass } \text{NaOCl}}{1500.0 \text{ g}}$$

$$\text{mass } \text{NaOCl} = 54.3 \text{ g}$$

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Chemical Society (Great Britain)



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Clinical Laboratory Procedures B. Carver, United States. Department of the Air Force, 1977 Fundamentals of Engineering Thermodynamics Michael J. Moran, Howard N. Shapiro, Daisy D. Boettner, Margaret B. Bailey, 2010-12-07 This leading text in the field maintains its engaging readable style while presenting a broader range of applications that motivate engineers to learn the core thermodynamics concepts Two new coauthors help update the material and integrate engaging new problems Throughout the chapters they focus on the relevance of thermodynamics to modern engineering problems Many relevant engineering based situations are also presented to help engineers model and solve these problems **How to Prepare for and Pass a High School Equivalency Exam** I. Edward Friedman, 1966 **Boilermaker 1 & C** United States. Bureau of Naval Personnel, 1969 **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1903 Includes list of members 1882 1902 and proceedings of the annual meetings and various supplements

Analytical Methods for Food Safety by Mass Spectrometry Guo-Fang Pang, 2018-06-21 Analytical Methods for Food Safety by Mass Spectrometry Volume Two Veterinary Drugs systematically introduces the Pesticide and Veterinary Drug Multiresidues Analytical Methods with discussions on 69 veterinary drug multiresidues chromatographic MS analytical methods that are capable of detecting over 200 veterinary drugs and chemical residues of 20 categories such as agonists lactams aminoglycosides amphenicols anabolic hormone anabolic steroids avermectins benzimidazole cephalosporins glucocorticoid steroids macrolides nitrofurans nitroimidazoles NSAIDs polyether polypeptides progestagens pyrazolones quinolones quinoxalines sedatives sulfonamides synthetic estrogens tetracyclines thyreostats and other toxins in animal and poultry tissues aquatic products milk milk powders and bee products This valuable book can be used as reference for not only university students but also technical personnel of different specialties who are engaged with study and applications such as food safety agricultural environment protection pesticide development and utilization in scientific research units institutions and quality inspection organizations Provides the chromatographic MS analytical technique for over 1000 commonly used veterinary and pesticide residues Presents the determinations of over 60 chemicals over 10 categories in plant derived products fruits vegetables grains teas Chinese medicinal herbs edible fungus fruit and vegetable juices and fruit wines Includes sections on animal derived products animal tissues aquatic products raw milk and milk powders etc Covers the latest information on sophisticated pre treatment techniques with a single sample pre treatment and simultaneous detection by GC MS and LC MS

Journal of the American Chemical Society American Chemical Society, 1927 Proceedings of the Society are included in v 1 59 1879 1937 **The Physical Chemistry of Electrolytic Solutions** Herbert Spencer Harned, Benton Brooks Owen, 1943 Teachers' and Students' Manual of Arithmetic John Appleby Ferrell, 1898 Illustrated official journal (patents), 1929 **Journal - Chemical Society, London** Chemical Society (Great Britain), 1900 *Chemical Abstracts*, 1907 *Journal of the Chemical Society* Chemical Society (Great Britain), 1900 Quarterly Journal of the Chemical

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