

## Unit 9 (Chapter 14): Gas Laws

Upon completion of Unit 9, you should be able to:

- Describe gas pressure, temperature, and volume in terms of gas particles
- Explain why gases can be compressed
- Identify the factors that determine and can affect gas pressure
- State, explain, and use Boyle's Law, Charles' Law, Gay-Lussac's Law, Dalton's Law of Partial Pressures, and the Ideal Gas Law
- Explain what is meant by an ideal gas—what assumptions are made
- Complete calculations similar to those found in the Gas Laws Packet and those done in the lab.

### Important!

Gases change the amount of volume they take up when the pressure or temperature changes, so scientists have agreed on a "standard temperature and pressure" (STP) to simplify some calculations.

*STP = standard temperature and pressure. This is 0°C and 1 atm of pressure.*

Gas laws also frequently use the Kelvin temperature scale, which has a bottom value of 0. This is unlike Celsius and Fahrenheit which can both go below zero if it is cold enough. To convert to Kelvin do the following:

$$^{\circ}\text{C} + 273 = \text{Kelvin temperature}$$

$$\text{Kelvin temperature} - 273 = ^{\circ}\text{C}$$

# Chapter 14 The Gas Laws Answer Key

**L Towne**



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