

17.1 A Model for Reaction Rates

Chemical Kinetics – the study of how fast chemical reactions occur.

Reaction Rates

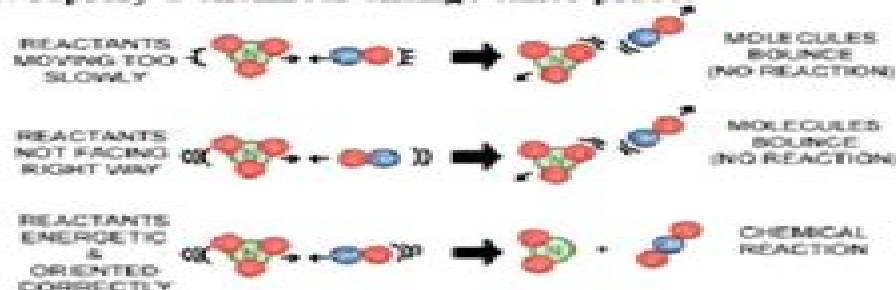
- Rate = change/unit of time
- Reaction rate = rate at which reactants are turned into products. This is expressed by using the concentration (molarity) of reactants or products changed in a given amount of time (s).
- Units of reaction rate = M/s or $mol/L \cdot s$

$$\text{Average rate of reaction} = \frac{\Delta[\text{reactant or product}]}{\Delta \text{time}}$$

- $[] = M = \text{Molarity}$

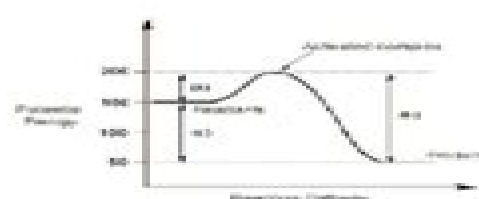
Collision theory – Particles must collide for chemical changes to occur.

- These particles may be ions, atoms or molecules.
- The collision of particles causes electrons to shift breaking old bonds and forming new ones.
- An **effective collision** is one in which the colliding particles approach each other at the proper orientation and with the proper energy to cause a reaction to take place.
- Thus, energy of collisions and the way in which particles collide affects how rapidly a chemical change takes place.



- **Energy in Chemical Reactions**

- Kinetic energy of the particles involved (kinetic energy is affected by the speed of the particle and/or its mass)
- **Activation Energy** = energy required to get a chemical reaction to proceed. This energy is above and beyond the energy initially held in the reactants.



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Carole H. McQuarrie



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