

Version A

Chapter 17 - Thermochemistry

Matching

Match each item with the correct statement below.

- | | |
|------------------|-------------|
| a. calorimeter | e. joule |
| b. specific heat | d. enthalpy |

- _____ 1. quantity of heat needed to raise the temperature of 1 g of a substance by 1°C
_____ 2. SI unit of energy
_____ 3. device used to measure the heat absorbed or released during a chemical or physical process
_____ 4. potential energy of the products minus potential energy of the reactants

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 5. A piece of metal is heated, then submerged in cool water. Which statement below describes what happens?
a. The temperature of the water will decrease and the temperature of the metal will increase.
b. The temperature of the water will increase and the temperature of the metal will stay the same.
c. The temperature of the water will decrease and the temperature of the metal will stay the same.
d. The temperature of the water will increase and the temperature of the metal will decrease.
- _____ 6. What would likely happen if you were to touch the flask in which an endothermic reaction were occurring?
a. The flask would probably feel cooler than before the reaction started.
b. The flask would probably feel warmer than before the reaction started.
c. The flask would feel the same as before the reaction started.
d. none of the above
- _____ 7. Which of the following is transferred due to a temperature difference?
a. chemical energy
b. potential energy
c. electrical energy
d. heat
- _____ 8. In an exothermic reaction, the energy stored in the chemical bonds of the reactants is _____.
a. equal to the energy stored in the bonds of the products
b. greater than the energy stored in the bonds of the products
c. less than the energy stored in the bonds of the products
d. less than the heat released
- _____ 9. A process that releases heat is a(n) _____.
a. endothermic process
b. polythermic process
c. exothermic process
d. isothermic process
- _____ 10. How many joules are in 148 calories? (1 cal = 4.18 J)
a. 6.61 J
b. 35.4 J
c. 148 J
d. 619 J

Chapter 17 Thermochemistry Test

LP Steffe



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