Chapter 19- Acids, Bases and Salts. GA chemistry standards: SC7.b: compare contrast and evaluate the nature of acids and bases SC7.b.1: Arrhenius, Bronsted-Lowery acid/bases SC7.b.2: strong vs. weak acids/bases in terms of percent dissociation SC7.b.3: hydronium ion concentration (hydrogen ion) SC7.b.4: Acid-Base neutralization

- 19.1 Acid-Base Theories
  - Properties of Acids
    - o Sour taste
    - React with metals to produce hydrogen gas
    - Aqueous solutions conduct electricity and are call electrolytes.
    - Will change the color of acid base indicator.
    - Law pH (below 7)
  - Properties of Bases
    - o Bitter taste
    - Feel slippery
    - Aqueous solutions conduct electricity and are call electrolytes
    - Will change the color of acid base indicator
    - o High pH (above 7)
  - Arrhenius Acids are defined as hydrogen-containing compounds that ionize to yield hydrogen ions (H\*\*) in aqueous solutions
    - Monoprotic acid are acids that contain ONE ionizable hydrogen. Examples: HCl or HNO.
    - Diprotic acid are acids that contain TWO ionizable hydrogen. Examples: H<sub>2</sub>S or H<sub>2</sub>SO<sub>x</sub>
    - Triprotic acids are acids that contain THREE ionizable hydrogen. Examples: H<sub>2</sub>PO<sub>4</sub>
  - Arrhenius bases are defined as compounds that ionize to yield hydroxide ions (OH\*) in aqueous solutions
  - Bronsted-Lowery defines an acid as a Hydrogen-lon donor
  - Bronsted-Lowery defines a base as a hydrogen-ion acceptor
    - Conjugate acid is the particle formed when a base gains a hydrogen ion
    - Conjugate base is the particle formed when an acid looses a hydrogen ion
    - Conjugate acid-base pair consist of two substances that are related by the loss or gain of a single hydrogen ion

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- Conjugate acid-base pairs: NH, &NH, \*, H,O & OH\*
- If the compounds differ by more that one hydrogen ion (or any other element) they CAN NOT be classified as conjugate acid-base pairs.
- Lewis acid is a substance that can accept a pair of electrons to form a covalent bond.
- Lewis base is a substance that can donate a pair of electrons to form a covalent bond.

Туре	Acid	Base
Arrhenius	H" producer (donor)	OH producer
Bronsted-Lowery	H <sup>*</sup> producer (donor)	H* acceptor
Lewis	Electron-pair acceptor	Electon-pair donor

- 19.2 Hydrogen lons and Acidity
  - Self-ionization is where water molecules production ions.

$$H_2O(I) \rightleftharpoons H^+(aq) + OH^-(aq)$$

- Hydrogen ion Hydroxide ion
- Neutral solution is where the [H\*\*] and [OH\*] are equal ( []-concentration of)

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C provides deep insights into different although comparable approaches which aim to frame to assess and to promote learning and learning outcomes in science education Smaller projects are presented as well as broad coordinated national programs. The papers in Part D outline the individual historical development from different national perspectives reflecting the deficits and problems that led to current reforms Finally a summary of the organizers analyses the conclusions from different vantage points. Basic Nursing Patricia Ann Potter, Anne Griffin Perry, 1991. Resources in Education, 1984.

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