

## Chapter 13: Properties of Solutions

Answers to selected problems:

23. a) unsaturated      b) saturated

24. a) 25 g

33. Solubility of Helium:  $5.6 \times 10^{-4} \text{ M}$ ; Solubility of  $\text{N}_2 = 9.0 \times 10^{-4} \text{ M}$

34. Solubility of Oxygen:  $2.5 \times 10^{-4} \text{ M}$ . Hint: Be sure to change torr to atm.

35.      a) 2.15 %;                      b) 3.15 ppm

38.      a) Mole fraction of phenol: 0.0285

         b) 5.66 %                      c) 0.638 m

40.      a) 0.125 M Aluminum Sulfate                      b) 0.140 M

         c) 0.63 M (hint: use  $MV = MV$ , see Chapter 4)

41.      a) 4.70 m                      b) 0.235 m

43.      \*\*To make this problem easier, I would assume 1.000 L or 1000 mL.

         a) 43.01%                      b) mole fraction = 0.122

         c) 7.69 m                      d) 5.828 M

46.      a) 0.0439                      b) 0.498 m                      c) 0.417 M

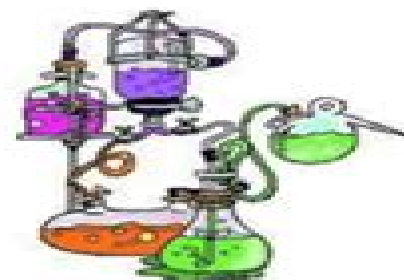
(Hint: on this problem, you will have to use the density of thiophene to get the mL of thiophene and you will also have to use the density of toluene to get the mass(g) of toluene.

48.      a) 0.278 mol                      b)  $6.25 \times 10^{-5}$  moles (assume the mass of the solvent is approx. equal to the mass of the solution for this problem)

         c)  $3.29 \times 10^{-3}$  mol

51.      \*\* Assume you have 1.000L or 1000 mL. You can also use density to get the total grams.

Answer: 71%



# Chapter 13 Review Solutions Section Answers

**Léo Pomar**



## Chapter 13 Review Solutions Section Answers:

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