

CHAPTER 15

ACIDS AND BASES

- 15.3** Table 15.2 of the text contains a list of important Brønsted acids and bases. (a) both (why?), (b) base, (c) acid, (d) base, (e) acid, (f) base, (g) base, (h) base, (i) acid, (j) acid.

- 15.4** Recall that the conjugate base of a Brønsted acid is the species that remains when one proton has been removed from the acid.

- (a) nitrite ion: NO_2^-
 (b) hydrogen sulfate ion (also called bisulfate ion): HSO_4^-
 (c) hydrogen sulfide ion (also called bisulfide ion): HS^-
 (d) cyanide ion: CN^-
 (e) formate ion: HCOO^-

- 15.5** In general the components of the conjugate acid–base pair are on opposite sides of the reaction arrow. The base always has one fewer proton than the acid.

- (a) The conjugate acid–base pairs are (1) HCN (acid) and CN^- (base) and (2) CH_3COO^- (base) and CH_3COOH (acid).
- (b) (1) HCO_3^- (acid) and CO_3^{2-} (base) and (2) HCO_3^- (base) and H_2CO_3 (acid).
- (c) (1) H_2PO_4^- (acid) and HPO_4^{2-} (base) and (2) NH_3 (base) and NH_4^+ (acid).
- (d) (1) HCHO (acid) and CHO^- (base) and (2) CH_3NH_2 (base) and CH_3NH_3^+ (acid).
- (e) (1) H_2O (acid) and OH^- (base) and (2) CO_3^{2-} (base) and HCO_3^- (acid).

- 15.6** The conjugate acid of any base is just the base with a proton added.

- (a) H_2S (b) H_2CO_3 (c) HCO_3^- (d) H_3PO_4 (e) H_2PO_4^-
(f) HPO_4^{2-} (g) H_2SO_4 (h) HSO_4^- (i) HSO_3^-

- 15.7** (a) The Lewis structures are



- (b) H^+ and $\text{C}_2\text{H}_2\text{O}_4$ can act only as acids, C_2HO_4^- can act as both an acid and a base, and $\text{C}_2\text{O}_4^{2-}$ can act only as a base.

- 15.8** The conjugate base of any acid is simply the acid minus one proton.

- (a) $\text{CH}_3\text{ClCOO}^-$ (b) IO_4^- (c) H_2PO_4^- (d) HPO_4^{2-} (e) PO_4^{3-}
(f) HSO_4^- (g) SO_4^{2-} (h) IO_3^- (i) SO_3^{2-} (j) NH_3
(k) HS^- (l) S^{2-} (m) OCl^-

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