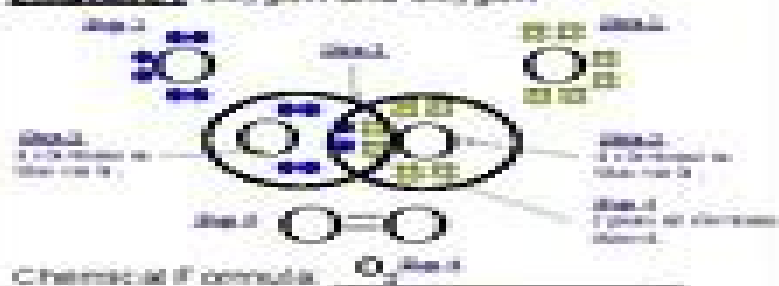


Covalent Bonding Practice

Instructions: Show how the following non-metals will covalently bond using the steps below.

1. Draw the Lewis dot structure using the number of valence electrons for each element. *Hint: You might want to use different colored pencils or dots and write the different elements in large numbers so which electrons belong to which element.*
2. Rearrange the electrons in between the atoms so each can fulfill the octet rule. *Hint: Hydrogen is an exception to the octet rule as it only needs 2 valence electrons to fulfill the outer shell (the noble electron configuration of Helium).*
3. Circle each atom while making sure there are 8 electrons (or 2 for hydrogen) in each circle, which fulfills the octet rule. Circles will overlap.
4. Count the pairs of electrons that are shared in between each atom in the overlapping circles. There may be one pair, two pairs, three pairs or four pairs that are shared.
5. Redraw each atom and show a line for each pair of electrons that are shared between the atoms.
6. Write the chemical formula using subscripts to indicate how many atoms there are of each element. *Hint: There is no need to write a subscript if there is only one of an element (oxygen). There are two Hydrogens and one oxygen in H_2O .*
7. Write the number of each type of covalent bond that is the atom makes depending on the number of electrons shared between each atom (1 pair=single, 2 pairs=double, 3 pairs=triple, 4 pairs=quadruple)

Example: Oxygen and Oxygen



Type of Bond: Double Bond

2. Hydrogen and Hydrogen

Chemical Formula: _____

Type of Bond: _____

3. Chlorine and Chlorine

Chemical Formula: _____

Type of Bond: _____

4. Nitrogen and Nitrogen

Chemical Formula: _____

Type of Bond: _____

Chapter 16 Covalent Bonding Worksheet Key

McGraw-Hill Staff



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Holt Science and Technology Holt, Rinehart and Winston Staff, 2001 **Books in Print Supplement**, 2002

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