

Acid/Base Titration – pH Calculations

Strong acid titrated with strong base



- Issue is which is in excess, (XS), H^+ or OH^-

- **Pre-equivalence strategy:**

1. H_3O^+ is in XS, not completely neutralized.
2. $(M_a V_a - M_b V_b) / (V_a + V_b) = [\text{unreacted acid}]$
3. $[\text{unreacted acid}] = [\text{H}^+]$
4. Calculate pH

- **Equivalence** pH = 7.00

- **Post-equivalence strategy:**

1. $(M_b V_b - M_a V_a) / (V_a + V_b) = [\text{XS base}]$
2. pH = pK_a + pOH

Weak acid titrated with a strong base

- To calculate pH we must now look at the conjugate base as it hydrolyzes water molecules.

- **Three different calculations:**

- 1 **pre-equivalence strategy:**

- a) determine moles /molarity of unreacted acid remaining;

initial moles acid - moles base added = moles acid left

$$(M_a V_a - M_b V_b) / (V_a + V_b) = [\text{unreacted acid}]$$

- b) determine moles/molarity of conjugate base released;

moles base added = moles conjugate released

$$(M_b V_b) / (V_a + V_b) = [\text{conjugate}]$$

- c) use K_a of the acid and calculate $[\text{H}^+]$ & pH.

Chapter 16 Review Acid Base Titration Ph Answers

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Yuan, Heddwen L. Brooks**



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