# Acids, Bases, and Salts Review

			ny compound dissociating s. Example: HNO,(aq.) →	
	solution to form ions. Example: $HNO_3(aq.) \rightarrow H^*(aq.) + NO_3(aq.)$ . Complete the following equation:			
	HCI (aq.) →			
@	An Arrhenius base is defined as any compound dissociating in an aqueous solution to form ions. Example: KOH (aq) → K <sup>*</sup> (aq) + OH <sup>*</sup> (aq). Complete the following equation:  NaOH (aq.) →			
3	Salts are compounds that dissociate in an aqueous solution releasing neither nor ions. Example: KCl (aq.) → K* (aq.) + Cl* (aq.).			
	Complete the following equation:			
		NaCl (aq.) →	<u> </u>	
<b>④</b>	A solution that has strong conductivity and reacts with the metal to make hydrogen gas is a			
	Strong Base	<b>b</b> Weak Base	Strong Acid	d Weak Acid
	Answers			
3	Which of the following is slippery to touch?			
	Strong Acid	<b>b</b> Strong Base	Strong Electrolyte	d Water
	Answers			
•	The hydronium ion is			
	■ H <sub>3</sub> O.	B H₂O	C OH-	d H <sub>2</sub>
	Answers			
0	A substance that changes color in an acidic or basic solution is called			
	<ul><li>Neutralizer</li></ul>	<b>b</b> Inhibitor	pH indicator	d Catalyst
	Answers			
(8)	Name the following acids and bases			
	HF		HNO3	
	кон		Mg(OH),	

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Holt Rinehart & Winston, Holt, Rinehart and Winston Staff

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