

## Practice Problems: Redox Reactions (Answer Key)

1. Determine the oxidation number of the elements in each of the following compounds:



**H: +1, O: -2, C: +4**



**N: 0**



**Zn: 2+, H: +1, O: -2**



**N: +3, O: -2**



**Li: +1, H: -1**



**Fe: +8/3, O: -2**

2. Identify the species being oxidized and reduced in each of the following reactions:



**$\text{Cr}^+$ : oxidized,  $\text{Sn}^{4+}$ : reduced**



**$\text{Hg}^{2+}$ : reduced, Fe: oxidized**



**As: oxidized,  $\text{Cl}_2$ : reduced**

3. Would you use an oxidizing agent or reducing agent in order for the following reactions to occur?

# Chapter 19 Practice Problems Redox Reactions

**Mark Anestis, Kellie Ploeger Cox**



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