

## Assessment

**Chapter Test B****Chapter: Acid-Base Titration and pH**

**PART I** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The pH scale generally ranges from  
a. 0 to 1.  
b. -1 to 1.  
c. 0 to 7.  
d. 0 to 14.
- \_\_\_\_\_ 2. During the titration of HCl with NaOH, a very rapid change in pH occurs  
a. when the first addition of known solution is made.  
b. when roughly equivalent amounts of  $\text{H}_3\text{O}^+$  ions and  $\text{OH}^-$  ions become present.  
c. at several points.  
d. at no point.
- \_\_\_\_\_ 3. A water solution is neutral if  
a. it contains no  $\text{H}_3\text{O}^+$  ions.  
b. it contains no ionized water molecules.  
c. it contains no  $\text{H}_3\text{O}^+$  ions or  $\text{OH}^-$  ions.  
d. the concentrations of  $\text{H}_3\text{O}^+$  ions and  $\text{OH}^-$  ions are equal.
- \_\_\_\_\_ 4. The antilog of a number  $N$  is  
a. the inverse of  $N$ .  
b. the square root of  $N$ .  
c. 10 raised to the power of  $N$ .  
d.  $N$  raised to the 10th power.
- \_\_\_\_\_ 5. Universal indicators  
a. are mixtures of several indicator solutions.  
b. are pure substances.  
c. have very brief color-change intervals.  
d. work well only for acidic solutions.
- \_\_\_\_\_ 6. A useful pH range for an indicator in neutralizations involving strong acids and weak bases is  
a. 1.2 to 3.0.  
b. 3.1 to 4.6.  
c. 6.0 to 7.6.  
d. 9.5 to 11.0.

# Chapter 1 acid Base Titration Ph Practice Test

**Neil D. Jespersen, Pamela Kerrigan**



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## **Table of Contents Chapter 1acid Base Titration Ph Practice Test**

1. Understanding the eBook Chapter 1acid Base Titration Ph Practice Test
  - The Rise of Digital Reading Chapter 1acid Base Titration Ph Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1acid Base Titration Ph Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 1acid Base Titration Ph Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1acid Base Titration Ph Practice Test
  - Personalized Recommendations
  - Chapter 1acid Base Titration Ph Practice Test User Reviews and Ratings
  - Chapter 1acid Base Titration Ph Practice Test and Bestseller Lists
5. Accessing Chapter 1acid Base Titration Ph Practice Test Free and Paid eBooks
  - Chapter 1acid Base Titration Ph Practice Test Public Domain eBooks
  - Chapter 1acid Base Titration Ph Practice Test eBook Subscription Services
  - Chapter 1acid Base Titration Ph Practice Test Budget-Friendly Options

6. Navigating Chapter 1acid Base Titration Ph Practice Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 1acid Base Titration Ph Practice Test Compatibility with Devices
  - Chapter 1acid Base Titration Ph Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 1acid Base Titration Ph Practice Test
  - Highlighting and Note-Taking Chapter 1acid Base Titration Ph Practice Test
  - Interactive Elements Chapter 1acid Base Titration Ph Practice Test
8. Staying Engaged with Chapter 1acid Base Titration Ph Practice Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 1acid Base Titration Ph Practice Test
9. Balancing eBooks and Physical Books Chapter 1acid Base Titration Ph Practice Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 1acid Base Titration Ph Practice Test
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 1acid Base Titration Ph Practice Test
  - Setting Reading Goals Chapter 1acid Base Titration Ph Practice Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 1acid Base Titration Ph Practice Test
  - Fact-Checking eBook Content of Chapter 1acid Base Titration Ph Practice Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements



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