

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a <u>proton donor</u>	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid	HCl	MINERAL ACIDS	They are derived from mineral salts i.e chlorides for HCl, sulphates for H_2SO_4 and nitrates for HNO_3
Sulphuric acid	H_2SO_4		
Nitric acid	HNO_3		

Other mineral acids known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H_2SO_3	Derived from sulphites	Ethanoic acid (CH_3COOH)	CITRIC ACID from lemons
Carbonic acid	H_2CO_3	Derived from carbonates		TARTARIC ACID from grapes
Phosphoric acid	H_3PO_4	Derived from phosphates	Methanoic acid ($HCOOH$)	ACETIC ACID from vinegar
Nitrous acid	HNO_2	Derived from nitrites		LACTIC ACID from sour milk
				Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> .
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Chapter 1acids Bases Study Guide

Susan McMurry



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Table of Contents Chapter 1acids Bases Study Guide

1. Understanding the eBook Chapter 1acids Bases Study Guide
 - The Rise of Digital Reading Chapter 1acids Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1acids Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1acids Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1acids Bases Study Guide
 - Personalized Recommendations
 - Chapter 1acids Bases Study Guide User Reviews and Ratings
 - Chapter 1acids Bases Study Guide and Bestseller Lists
5. Accessing Chapter 1acids Bases Study Guide Free and Paid eBooks
 - Chapter 1acids Bases Study Guide Public Domain eBooks
 - Chapter 1acids Bases Study Guide eBook Subscription Services
 - Chapter 1acids Bases Study Guide Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Chapter 1acids Bases Study Guide Compatibility with Devices
 - Chapter 1acids Bases Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1acids Bases Study Guide
 - Highlighting and Note-Taking Chapter 1acids Bases Study Guide
 - Interactive Elements Chapter 1acids Bases Study Guide
8. Staying Engaged with Chapter 1acids Bases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1acids Bases Study Guide
9. Balancing eBooks and Physical Books Chapter 1acids Bases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1acids Bases Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 1acids Bases Study Guide
 - Setting Reading Goals Chapter 1acids Bases Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 1acids Bases Study Guide
 - Fact-Checking eBook Content of Chapter 1acids Bases Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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