
ANATOMY & PHYSIOLOGY

CHAPTER 2 STUDY GUIDE

WEEK 1

Chemistry of Life

Levels of Chemical Organization

- **Atoms:** The basic unit of matter, containing a nucleus made up of protons (positively charged) and neutrons (uncharged).
 - **Atomic Number:** Number of protons in the nucleus; defines the element.
 - **Atomic Mass:** Total number of protons and neutrons.
 - **Electrons:** Negatively charged particles that orbit the nucleus in energy levels (orbitals), each capable of holding up to eight electrons.
-

Elements, Molecules, and Compounds

- **Element:** A pure substance made up of only one kind of atom.
 - **Molecule:** A group of atoms bonded together.
 - **Compound:** A substance with molecules that contains more than one kind of atom.
-

Chemical Bonding

- **Purpose:** Bonds form to stabilize atoms by filling their outer energy levels.
 - **Types of Bonds:**
 - **Ionic Bonds:** Formed by the attraction of oppositely charged ions.
 - **Positive Ion:** Atom that loses electrons (e.g., Na^+ , Ca^{2+}).
 - **Negative Ion:** Atom that gains electrons (e.g., Cl^-).
 - **Covalent Bonds:** Formed by the sharing of electrons between atoms; generally stable and do not easily dissociate in water.
 - **Hydrogen Bonds:** Weak attractions that hold molecules together in folded shapes or groups; do not create new molecules.
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Chapter 12 Blood Study Guide

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