

Name _____ Date _____ Class _____

CHAPTER 16

STUDY GUIDE FOR CONTENT MASTERY

Energy and Chemical Change

Section 16.1 Energy

In your notebook, read about the nature of energy.

In the space at the left, write true if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- True* _____ 1. Energy is the ability to do work or produce heat.
- cannot be* _____ 2. The law of conservation of energy states that energy can be created and destroyed.
- True* _____ 3. Chemical potential energy is energy stored in a substance because of its composition.
- True* _____ 4. Heat is a form of energy that flows from a warmer object to a colder object.
- True* _____ 5. A calorie is the amount of energy required to raise the temperature of one gram of pure water by one degree Celsius.
- False* _____ 6. A calorie is the SI unit of heat and energy.
- True* _____ 7. The specific heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.

True _____ 8. Kinetic energy is energy of motion.

both potential and

kinetic energy

_____ 9. Chemical potential energy is a chemical reaction *contains only* potential energy.

1000 calories _____ 10. One nutritional Calorie is equal to 100 calories.

True _____ 11. One calorie equals 4.184 joules.*True* _____ 12. When a fuel is burned, some of its chemical potential energy is lost as heat.*multiply* _____ 13. To convert kilojoules to joules, divide the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following questions. Show all your work.

14. If the temperature of a 500-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/g·°C.

$$q = c \times m \times \Delta T$$

$$q = 4.184 \text{ J/g}\cdot^{\circ}\text{C} \times 500 \text{ g} \times 2.00^{\circ}\text{C}$$

$$q = 4.184 \times 10^3 \text{ J} = 4.18 \text{ kJ}$$

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CHAPTER 16

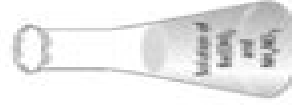
STUDY GUIDE FOR CONTENT MASTERY

Section 16.2 Heat in Chemical Reactions and Processes

In your notebook, read about measuring heat and about chemical energy and the universe.

For each item in Column A, write the letter of the matching item in Column B.

- | Column A | Column B |
|---|--------------------------------|
| b. _____ 1. An insulated device used to measure the amount of heat absorbed or released during a chemical or physical process | a. system |
| c. _____ 2. The study of heat changes that accompany chemical reactions and phase changes | b. calorimeter |
| d. _____ 3. The specific part of the universe that contains the reaction or process you wish to study | c. thermodynamics |
| e. _____ 4. The change in enthalpy in a chemical reaction | d. universe |
| f. _____ 5. A system plus its surroundings | e. enthalpy |
| g. _____ 6. The heat content of a system at constant pressure | f. enthalpy change of reaction |
| h. _____ 7. Everything in the universe except the system being studied | g. surroundings |



Use the illustration to answer the following questions.

8. A scientist is studying the solution in the flask. What is the system?
The system is the solution of NaOH and Na₃PO₄.
9. What are the surroundings?
The surroundings include everything except the solution.
10. What is the universe?
The universe is the solution plus the surroundings.

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Chapter 16 Study Guide Answers Energy Chemical Change

William H. Myers, William T. Mooney



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