

Energy and Chemical Change

Section 16.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- | | |
|-------|---|
| _____ | 1. <i>Energy</i> is the ability to do work or produce heat. |
| _____ | 2. The law of conservation of energy states that energy <i>can be</i> created and destroyed. |
| _____ | 3. Chemical potential energy is energy stored in a substance because of its <i>composition</i> . |
| _____ | 4. <i>Heat</i> is a form of energy that flows from a warmer object to a cooler object. |
| _____ | 5. A calorie is the amount of energy required to raise the temperature of <i>one gram</i> of pure water by one degree Celsius. |
| _____ | 6. A <i>calorie</i> is the SI unit of heat and energy. |
| _____ | 7. The <i>specific heat</i> of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius. |
| _____ | 8. <i>Kinetic energy</i> is energy of motion. |
| _____ | 9. Chemicals participating in a chemical reaction contain <i>only potential energy</i> . |
| _____ | 10. One nutritional Calorie is equal to <i>100 calories</i> . |
| _____ | 11. One calorie equals <i>4.184 joules</i> . |
| _____ | 12. When a fuel is burned, some of its <i>chemical potential energy</i> is lost as heat. |
| _____ | 13. To convert kilojoules to joules, <i>divide</i> the number of kilojoules by 1000 joules/1 kilojoule. |

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

Chapter 16 Study Guide For Content Mastery Answers Chemistry

Steven S. Zumdahl, Paul B. Kelter



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Chemistry Steven S. Zumdahl, Susan A. Zumdahl, 2000 Zumdahl's conceptual approach to problem solving early and thorough treatment of reactions and integration of descriptive chemistry and chemical principles throughout are key elements to the success of this bestseller By encouraging students to understand qualitative concepts before considering quantitative problems Chemistry helps students build a solid understanding of topics and avoid common misconceptions The text also includes new Chemical Impact boxes that present everyday applications of chemistry plus over 200 redesigned illustrations with greater clarity and emphasis on molecular level illustrations **Resources in Education**, 1984

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Essentials of Ecology* George Tyler Miller, 2006-09 G Tyler Miller's worldwide bestsellers have evolved right along with the changing needs of your diverse student population Focused specifically on energizing and engaging all your students Miller and new contributor Scott Spoolman have been at work scrutinizing every line enhancing clarifying and streamlining to reduce word density as well as updating with the very latest environmental news and research The resulting texts are shorter clearer and so engaging that your students will actually want to read their assignments The ideal alternative to ecology texts that tend to be too difficult for non majors this succinct 13 chapter full color textbook covers scientific principles and concepts ecosystems evolution biodiversity population ecology and more New to this edition for instructors is PowerLecture a one stop shop for lecture prep that includes everything you need to create dynamic lectures all in one place **Documentation Abstracts**, 1985

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Forthcoming Books Rose Arny, 1997-04 **International Books in Print**, 1987 **Glencoe Chemistry** Glencoe/McGraw-Hill, 2002 **Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition** McGraw Hill, 2001-03-30 Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats There are worksheets and study tools for each section of the text that help teachers track students progress toward understanding concepts Guided Reading Activities help students identify and comprehend the important information in each chapter *Study Guide for Chemical Principles [by] Steven S. Zumdahl* Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples [Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry](#) Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and

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