

Name: _____ Course and Section #: _____

Chapter 3: Acids, Bases, pH and Buffer Worksheet (15pts)

1. (1pt) H⁺ is the _____ ion.
2. (1pt) OH⁻ is the _____ ion.
3. (3pts) You have a solution of pH 3:
 1. The hydrogen ion concentration is _____.
 2. The hydroxide ion concentration is _____.
 3. Is the solution an acid or base? _____.
4. (2pts) The pH of a solution with a hydrogen ion concentration of 10^{-9} is
 1. _____.
 2. Is this solution an acid or a base? _____.
5. (2pts) The pH of a solution with a hydroxide ion concentration of 10^{-9} is
 1. _____.
 2. Is this solution an acid or a base? _____.
6. (3pts) You change the pH of a solution from 9 to 11.
 1. By what fold did your H⁺ concentration change? (circle one)
a. 2-fold b. 20-fold c. 10-fold d. 100-fold e. 1.2-fold
 2. Did the H⁺ concentration increase or decrease as you went from pH 9 to pH 11?
 3. Is the final solution with a pH of 11 more acidic or more basic than the initial solution with a pH of 9?
7. (1pt) What is a buffer?
8. (1pt) If you have a buffered solution of pH 6 and you add NaOH to the solution (a base) what is the most likely pH after the addition of the NaOH.
a. 11.4 b. 6.3 c. 2.1
9. (1pt) If you have a non-buffered solution of pH 7 and you add HCl to the solution (an acid) what is the most likely pH reading after the addition of the HCl. (1pt)
a. 1.2 b. 6.9 c. 10.7

Chapter 17 Reinforcement Acids Bases Answers

Jianjun Gao



Chapter 17 Reinforcement Acids Bases Answers:

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