

$$86.2^{\circ}\text{C} = T_f$$

8. For the combustion of methane, calculate the enthalpy change for the burning of 10.00 g of methane.



$$\Delta H = 10.00 \text{ g CH}_4 \times \frac{1 \text{ mol CH}_4}{16.05 \text{ g CH}_4} \times \frac{-890 \text{ kJ}}{1 \text{ mol CH}_4} =$$

$$\Delta H = -555 \text{ kJ}$$

17.3 Heat in Changes of State

9. Read heating curves (See *Changes of State* w/s) Label the phases, boiling point and melting point here:

boiling point = 20°C , melting point = 10°C



a = solid
b = melting
c = liquid
d = boiling
e = gas

Indicate whether a heating curve would be *flat* or *rising* in 10-14.

- | | |
|--------------------------------------|--------------------------------------|
| 10. liquid is boiling - flat | 13. PE is increasing - flat |
| 11. solid is warming - rising | 14. KE is increasing - rising |
| 12. solid is melting - flat | |

Heat of Fusion and Heat of Vaporization Problems

Values for Water: (SHOW YOUR WORK FOR FULL CREDIT!!)

$$\Delta H_{\text{fus}} = 6.02 \text{ kJ/mol}$$

$$\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$$

$$C_{\text{H}_2\text{O}(s)} = 2.03 \text{ J/g}^{\circ}\text{C}$$

$$C_{\text{H}_2\text{O}(l)} = 4.184 \text{ J/g}^{\circ}\text{C}$$

$$C_{\text{H}_2\text{O}(g)} = 2.02 \text{ J/g}^{\circ}\text{C}$$

15. How much heat needs to be added to 10090 grams of ice at -10.0°C so that it all becomes a vapor at 120.0°C ?

$$\text{A } q = mC\Delta T = (10090\text{g})(2.03 \text{ J/g}^{\circ}\text{C})(0 - (-10.0)) = 204827 \text{ J}$$

$$\text{B } \Delta H = \frac{10090 \text{ g H}_2\text{O}}{1} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{6.02 \text{ kJ}}{1 \text{ mol H}_2\text{O}} \times \frac{1000 \text{ J}}{1 \text{ kJ}} = 3370799.1 \text{ J}$$

$$\text{C } q = mC\Delta T = (10090\text{g})(4.184 \text{ J/g}^{\circ}\text{C})(100-0) = 4221656 \text{ J}$$

$$\text{D } \Delta H = \frac{10090 \text{ g H}_2\text{O}}{1} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{40.7 \text{ kJ}}{1 \text{ mol H}_2\text{O}} \times \frac{1000 \text{ J}}{1 \text{ kJ}} = 22789289.7 \text{ J}$$

$$\text{E } q = mC\Delta T = (10090\text{g})(2.02 \text{ J/g}^{\circ}\text{C})(120-100) = 407636 \text{ J}$$

Chapter 17 Thermochemistry Interpreting Graphics

Answers

Mark Steyvers



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