8. For the combustion of methane, calculate the enthalpy change for the burning of 10.00 g of methane.

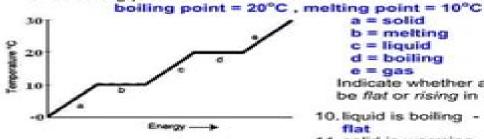
$$CH_4(g) + 2O_2(g) \rightarrow CO_2(g) + 2H_2O(I) + 890 \text{ kJ}$$

 $\Delta H = 10.00gCH_4 \times \frac{1moICH_4}{16.05gCH_4} \times \frac{-890 \text{ kJ}}{1moICH_4} =$

$$\Delta H = -555 \text{ kJ}$$

17.3 Heat in Changes of State

Read heating curves (See Changes of State w/s) Label the phases, boiling point and melting point here:



a = solidb = melting c = liquid d = boiling e = gas

Indicate whether a heating curve would be flat or rising in 10-14.

 liquid is boiling - 13.PE is increasing - flat flat 11. solid is warming — 14. KE is increasing —

rising

rising 12. solid is melting -

Heat of Fusion and Heat of flat Vaporization Problems Values for Water: (SHOW YOUR WORK FOR FULL CREDIT!!) $\Delta H_{tus} = 6.02 \text{ kJ/mol}$ $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$ CH2O(1) = 4.184 J/g°C CH2O(1) = 2.02 J/g°C CH20(s) = 2.03 J/g°C

15. How much heat needs to be added to 10090 grams of ice at -10.0°C so that it all becomes a vapor at 120.0°C?

$$\mathbf{A} = mC\Delta T = (10090g)(2.03 \text{ J/g}^{\circ}\text{C})(0-(-10.0)) = 204827 \text{ J}$$

$$\mathbf{B} \ \Delta H = \frac{10090 \ g \ H_{2}O}{I} \times \frac{1 \ mol \ H_{2}O}{I8.02 \ g \ H_{2}O} \times \frac{6.02 \ kJ}{1 \ mol \ H_{2}O} \times \frac{1000 \ J}{1 kJ} = \mathbf{3370799.1} \ \mathbf{J}$$

$$\mathbf{D} \ \Delta H = \frac{10090 \ g \ H_{2}O}{I} \times \frac{1 \ mol \ H_{2}O}{I8.02 \ g \ H_{2}O} \times \frac{40.7 kJ}{1 \ mol \ H_{2}O} \times \frac{1000 \ J}{1 kJ} = \mathbf{22789289.7 \ J}$$

$$\mathbf{E}_{q} = mC\Delta T = (10090g)(2.02 \text{ J/g}^{\circ}\text{C})(120-100) = 407636 \text{ J}$$

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Mark Steyvers

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