

## Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of  $\text{C}_6\text{H}_{12}\text{O}_6$ .

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia ( $\text{NH}_3$ ) can be produced?

- 3) Assume that in the decomposition of potassium chlorate,  $\text{KClO}_3$ ,



80.5 moles of  $\text{O}_2$  form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin"  $\text{PtCl}_2(\text{NH}_3)_2$ , has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g  $\text{K}_2\text{PtCl}_4$ ?

B. How many moles of  $\text{NH}_3$  would be needed if 3.8 moles of  $\text{K}_2\text{PtCl}_4$  were present?

# Chapter 12 Stoichiometry Work Answers

**Ray Boyington, William L. Masterton**



## Chapter 12 Stoichiometry Work Answers:

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*Student's Guide to Masterton and Slowinski's Chemical Principles* Ray Boyington, William L. Masterton, 1977      *Three Cognitive Skills in Chemistry and Their Application to Stoichiometry* Ardas Ozsogomonyan, 1977      Chemistry David W. Oxtoby, Wade A. Freeman, Toby F. Block, 1998 Textbook of chemistry concepts      **CHEMISTRY** SILBERBERG, 2003      *Engineering and Cement World* , 1928      **Holt Chemistry** Ralph Thomas Myers, 2004      **Chemistry** John S. Phillips, Cheryl Wistrom, 2000

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## **Table of Contents Chapter 12 Stoichiometry Work Answers**

1. Understanding the eBook Chapter 12 Stoichiometry Work Answers
  - The Rise of Digital Reading Chapter 12 Stoichiometry Work Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 12 Stoichiometry Work Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 12 Stoichiometry Work Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 12 Stoichiometry Work Answers
  - Personalized Recommendations
  - Chapter 12 Stoichiometry Work Answers User Reviews and Ratings
  - Chapter 12 Stoichiometry Work Answers and Bestseller Lists
5. Accessing Chapter 12 Stoichiometry Work Answers Free and Paid eBooks
  - Chapter 12 Stoichiometry Work Answers Public Domain eBooks
  - Chapter 12 Stoichiometry Work Answers eBook Subscription Services
  - Chapter 12 Stoichiometry Work Answers Budget-Friendly Options
6. Navigating Chapter 12 Stoichiometry Work Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 12 Stoichiometry Work Answers Compatibility with Devices
  - Chapter 12 Stoichiometry Work Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 12 Stoichiometry Work Answers
  - Highlighting and Note-Taking Chapter 12 Stoichiometry Work Answers
  - Interactive Elements Chapter 12 Stoichiometry Work Answers
8. Staying Engaged with Chapter 12 Stoichiometry Work Answers



- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chapter 12 Stoichiometry Work Answers
- 9. Balancing eBooks and Physical Books Chapter 12 Stoichiometry Work Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 12 Stoichiometry Work Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 12 Stoichiometry Work Answers
  - Setting Reading Goals Chapter 12 Stoichiometry Work Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 12 Stoichiometry Work Answers
  - Fact-Checking eBook Content of Chapter 12 Stoichiometry Work Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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