#### CHAPTER 13

### 13.1 States of matter review

Gases: contain particles that are far apart in random motion, and independent of one another.

- Little contact
- No attractive forces
- Random motion

**Solids**: contain particles very close in space and maintain a rigid shape. Significant attractive forces exist between particles.

- Close contact
- Strong attractive forces
- Rigid shape

Liquids: intermediate between gases and solids.

Contain particles close to one another but have fluidity (can assume the shape of a container).

Significant attractive forces exist between particles in a liquid.

- Close contact
- Some attractive forces.
- Fluid shape

### 13.2 Properties of liquids

### Evaporation

### Evaporation or Vaporization:

Escape of molecules from the liquid to the gas phase.

Liquid Vapor

Kinetic Energy: Energy associated with movement.

- Molecules in the liquid state have different kinetic energies (KE's).
- Those with higher KE's can overcome attractive forces between particles and escape to the gas phase.
- We can increase the kinetic energy (or movement of the molecules in liquid state by adding heat (boiling).

### Sublimation

Phase change from the solid to gas phase that bypasses the liquid state.

Solid vapor

Examples:

### Fusion.

Molecules can go from the liquid to solid phase.

This process is called "fusion"

Ex. Liquid solid

Under certain circumstances, gas can transform directly into a solid. This process is called "deposition".

Example: Water vapor to ice - Water vapor transforms directly into ice without becoming a liquid, a process that often occurs on windows during the winter months.

## Condensation

Molecules from the liquid phase can escape to the vapor phase through evaporation.

# **Chapter 13 Study Guide Gases Answers**

**Marty Machowski** 

## **Chapter 13 Study Guide Gases Answers:**

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