

Chapter 4 Practice Problems

1. How many grams of Li_3N can be formed from 1.75 moles of Li ? Assume an excess of nitrogen.



$$1.75 \text{ mol Li} \left(\frac{2 \text{ mol Li}_3\text{N}}{6 \text{ mol Li}} \right) \left(\frac{35.01 \text{ g Li}_3\text{N}}{1 \text{ mol}} \right) = \boxed{172.535 \text{ g Li}_3\text{N}}$$

2. How many moles of oxygen are formed when 58.6 g of KNO_3 decomposes according to the following reaction? The molar mass of KNO_3 is 101.11 g/mol.



$$58.6 \text{ g} \left(\frac{1 \text{ mol}}{101.11 \text{ g}} \right) \left(\frac{5 \text{ mol O}_2}{4 \text{ mol KNO}_3} \right) = \boxed{0.724 \text{ mol O}_2}$$

3. How many moles of nitrogen are formed when 58.6 g of KNO_3 decomposes according to the following reaction? The molar mass of KNO_3 is 101.11 g/mol.



$$58.6 \text{ g} \left(\frac{1 \text{ mol}}{101.11 \text{ g}} \right) \left(\frac{2 \text{ mol N}_2}{4 \text{ mol KNO}_3} \right) = \boxed{0.290 \text{ mol N}_2}$$

4. A 14.01 g sample of N_2 reacts with 3.02 g of H_2 to form ammonia (NH_3). If ammonia is the only product, what mass of ammonia is formed?

$$14.01 \text{ g} \left(\frac{1 \text{ mol}}{28.02} \right) \left(\frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \right) = 1.00 \text{ mol}$$

$$3.02 \text{ g} \left(\frac{1 \text{ mol}}{2.016} \right) \left(\frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) = 0.999 \text{ mol} \left(\frac{17.034}{1 \text{ mol}} \right) = \boxed{17.02 \text{ g NH}_3}$$

5. Consider the following balanced reaction. How many grams of water are required to form 75.9 g of HNO_3 ? Assume that there is excess NO_2 present. The molar masses are as follows: $\text{H}_2\text{O} = 18.02$ g/mol, $\text{HNO}_3 = 63.02$ g/mol.



$$75.9 \text{ g} \left(\frac{1 \text{ mol}}{63.02 \text{ g}} \right) \left(\frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol HNO}_3} \right) \left(\frac{18.02 \text{ g}}{1 \text{ mol}} \right) = \boxed{10.9 \text{ g H}_2\text{O}}$$

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