

**14****THE BEHAVIOR OF GASES****Chapter Test A****A. Matching**

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

**Column A****Column B**

- |  |  |
|--|--|
| _____ 1. ideal gas constant ( $R$ )        | a. The volume of a fixed mass of gas is directly proportional to the Kelvin temperature if the volume is kept constant.                                      |
| _____ 2. Boyle's law                       | b. At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases. |
| _____ 3. Dalton's law of partial pressures | c. $8.31 \frac{\text{L} \cdot \text{kPa}}{\text{K} \cdot \text{mol}}$  |
| _____ 4. ideal gas law                     | d. the contribution each gas in a mixture makes to the total pressure  |
| _____ 5. combined gas law                  | e. A gas tends to move to an area of lower concentration until the concentration is uniform throughout.  |
| _____ 6. Charles's law                     | f. $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$   |
| _____ 7. diffusion                         | g. $P \times V = n \times R \times T$  |
| _____ 8. partial pressure                  | h. For a given mass of gas at constant temperature, the volume of gas varies inversely with the pressure.  |

**B. Multiple Choice**

Choose the best answer and write its letter on the line.

- |  |                    |                  |
|--|--------------------|------------------|
| _____ 9. As the temperature of a fixed volume of gas increases, the pressure will  | a. vary inversely. | c. be unchanged. |
|  | b. decrease.       | d. increase.     |
| _____ 10. A breathing mixture used by deep-sea divers contains helium, oxygen, and carbon dioxide. What is the partial pressure of oxygen at 101.3 kPa total pressure if $P_{\text{He}} = 84.0 \text{ kPa}$ and $P_{\text{CO}_2} = 0.10 \text{ kPa}$ ? | a. 10.3 kPa        | c. 34.4 kPa      |
|  | b. 17.2 kPa        | d. 185.4 kPa     |

# Chapter 14 The Behavior Of Gases Chapter Test A Answers

**Linda D. Williams**



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