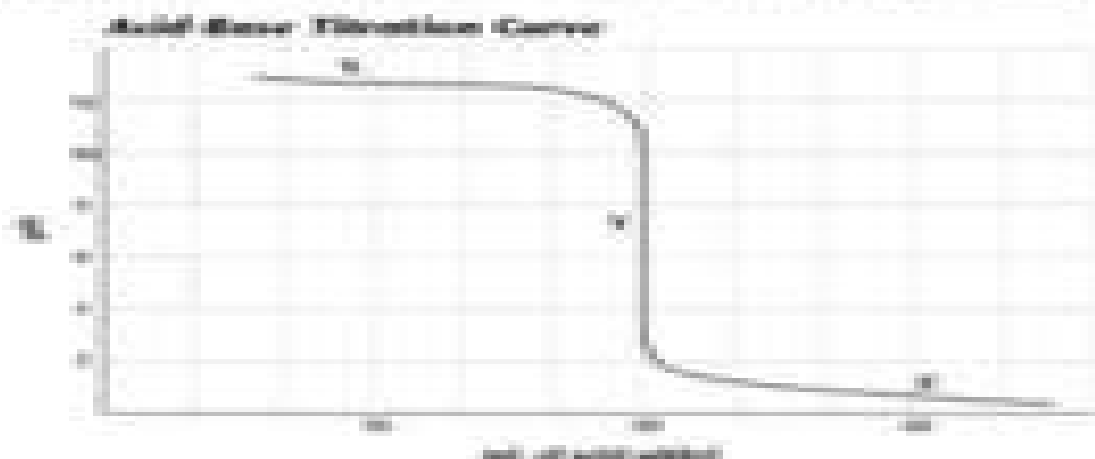


CHAPTER 13 REVIEW*Acid-Base Titration and pH***SECTION 2**

SHORT ANSWER Answer the following questions in the space provided.

5. Below is a pH curve from an acid-base titration. The x-axis labels three points: A, B, and C.



- _____ a. Which point represents the equivalence point?
- _____ b. At which point is there excess acid in the system?
- _____ c. At which point is there excess base in the system?
- _____ d. If the three solutions in (a), (b), and (c) were equal in volume of 100 mL each, how much of base, how many moles of H^+ , and how many moles of OH^- would be present in the solution?

PROBLEM Write the answer on the line to the left. Show all your work in the space provided.

6. A standardized solution of NaOH of 10.0 mL is mixed with a standard solution of unknown hydrochloric acid solution by combining, and the solution is taken 10.0 mL of the base for neutralization. What mL of the acid.
- a. Write the balanced molecular equation for this neutralization reaction.

Chapter 15 Acid Base Titration Ph Section 2 Answers

Mary Millhollon, Richard H. Langley



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