#### Version A

# Chapter 17 - Thermochemistry

#### Matching

		Match each item with the correct st a. calorimeter	atement below. c.	joule
		<ul> <li>specific heat</li> </ul>	d.	enthalpy
	1.	quantity of heat needed to raise the temperature of 1 g of a substance by 1°C		
	2.	SI unit of energy		
	3.	device used to measure the heat absorbed or released during a chemical or physical process		
	4.	potential energy of the products minus potential energy of the reactants		
		Choice		
Identi)	for their	letter of the choice that best comple	tes the statemer	it or answers the question.
	5.	A piece of metal is heated, then submerged in cool water. Which statement below describes what happens?  a. The temperature of the water will decrease and the temperature of the metal will increase.  b. The temperature of the water will increase and the temperature of the metal will stay the same.		
		<ul> <li>The temperature of the water w same.</li> </ul>	ill decrease and	the temperature of the metal will stay the
		d. The temperature of the water w	ill increase and	the temperature of the metal will decrease.
	6.			
		<ul> <li>The flask would probably feel of</li> </ul>		
		<ul> <li>The flask would probably feel v</li> </ul>		
		<ul> <li>The flask would feel the same a</li> <li>none of the above</li> </ul>	is before the rea	ction started.
	7.			
		<ul> <li>a. chemical energy</li> </ul>		electrical energy
		<ul> <li>b. potential energy</li> </ul>	d.	heat
	8.			
		equal to the energy stored in the bonds of the products     greater than the energy stored in the bonds of the products		
		<ul> <li>less than the energy stored in th</li> <li>d. less than the heat released</li> </ul>	e bonds of the j	products
		u. sess than the near reseased		
	9.	A process that releases heat is a(n)		
		a. endothermic process		exothermic process
		<ul> <li>b. polythermic process</li> </ul>	d.	isothermic process
	10.	How many joules are in 148 calories? (1 cal = 4.18 J)		
		a. 6.61 J	C.	148 J
		b. 35.4 J	d.	619 J

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**JA Banks** 

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