

Chapter 19- Acids, Bases and Salts. GA chemistry standards:
 SC7.b: compare contrast and evaluate the nature of acids and bases
 SC7.b.1: Arrhenius, Bronsted-Lowery acid/bases
 SC7.b.2: strong vs. weak acids/bases in terms of percent dissociation
 SC7.b.3: hydronium ion concentration (hydrogen ion)
 SC7.b.4: Acid-Base neutralization

19.1 Acid-Base Theories

- Properties of Acids
 - Sour taste
 - React with metals to produce hydrogen gas
 - Aqueous solutions conduct electricity and are call electrolytes
 - Will change the color of acid base indicator
 - Low pH (below 7)
- Properties of Bases
 - Bitter taste
 - Feel slippery
 - Aqueous solutions conduct electricity and are call electrolytes
 - Will change the color of acid base indicator
 - High pH (above 7)
- Arrhenius Acids are defined as hydrogen-containing compounds that ionize to yield hydrogen ions (H^+) in aqueous solutions
 - Monoprotic acid are acids that contain ONE ionizable hydrogen. Examples: HCl or HNO_3
 - Diprotic acid are acids that contain TWO ionizable hydrogen. Examples: H_2S or H_2SO_4
 - Triprotic acids are acids that contain THREE ionizable hydrogen. Examples: H_3PO_4
- Arrhenius bases are defined as compounds that ionize to yield hydroxide ions (OH^-) in aqueous solutions
- Bronsted-Lowery defines an acid as a Hydrogen-ion donor
- Bronsted-Lowery defines a base as a hydrogen-ion acceptor
 - Conjugate acid is the particle formed when a base gains a hydrogen ion
 - Conjugate base is the particle formed when an acid loses a hydrogen ion
 - Conjugate acid-base pair consist of two substances that are related by the loss or gain of a single hydrogen ion



- Conjugate acid-base pairs: NH_3 & NH_4^+ , H_2O & OH^-
 - If the compounds differ by more than one hydrogen ion (or any other element) they CAN NOT be classified as conjugate acid-base pairs.
 - Lewis acid is a substance that can accept a pair of electrons to form a covalent bond
 - Lewis base is a substance that can donate a pair of electrons to form a covalent bond
- | Type | Acid | Base |
|-----------------|-------------------------------|------------------------|
| Arrhenius | H^+ producer (donor) | OH^- producer |
| Bronsted-Lowery | H^+ producer (donor) | H^+ acceptor |
| Lewis | Electron-pair acceptor | Electron-pair donor |

19.2 Hydrogen Ions and Acidity

- Self-ionization is where water molecules production ions

$$\text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{OH}^-(\text{aq})$$

Hydrogen ion Hydroxide ion
- Neutral solution is where the $[\text{H}^+]$ and $[\text{OH}^-]$ are equal ([] = concentration of)

Chapter 19 Acids And Bases Study Guide

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Pediatric Critical Care Study Guide Steven E. Lucking, Frank A. Maffei, Robert F. Tamburro, Neal J. Thomas, 2012-04-10 This is the first comprehensive study guide covering all aspects of pediatric critical care medicine It fills a void that exists in learning resources currently available to pediatric critical care practitioners The major textbooks are excellent references but do not allow concise reading on specific topics and are not intended to act as both text and study guide There are also several handbooks available but these are usually written for general pediatric residents and lack the advanced physiology and pathophysiology required for the higher level pediatric critical care practitioner **Study Guide to Accompany Structure and Function of the Body** Linda Swisher, Gary A. Thibodeau, 2003-11 This comprehensive resource provides a variety of exercises for readers to apply and test their knowledge It contains matching fill in the blanks crossword puzzles word find unscramble the word application questions diagrams and page number references in the answer key December 2003 **Know Your 'O' Level Chemistry - A Study Guide , MCAT Complete; Medical College Admission Test Comprehensive Study Guide Physical Sciences; Biological Sciences; Verbal Reasoning** James Adams, Maryn Munley, Nathan Eisenbeis, Caroline Miller, Jesse Shetler, 2021-10-12 The MCAT The Medical College Admission Test MCAT is a standardized multiple choice examination designed to assess the examinee s problem solving critical thinking and knowledge of science concepts and principles prerequisite to the study of medicine Scores are reported in Physical Sciences Verbal Reasoning and Biological Sciences Study guide covers AAMC Association of American Medical Colleges content Physical Sciences Biological Sciences Verbal Reasoning Mathematics Concepts The Cell Chromosomes Reproduction Implantation Microorganisms Biochemistry Human Physiology The Heart The Lymphatic System GI Tract Musculoskeletal System Kidney Hormones Nerves Skin Genetics Populations and Evolution Elements Hund s Rule and Radiation The Periodic Table Covalent Bonds Molecular Shapes General Chemistry Kinetic Molecular Theory Phase Change Solutions Oxidation Numbers Entropy Acids and Bases Galvanic and Electrolytic Cells Carbon Stereochemistry Alkanes and Alkenes Hydrogen Bonding Alcohols Phenols Aldehydes and Ketones m Carboxylic Acids Ether Ammonia Amino Acids Carbohydrates Spectroscopy Separation and Distillation Vectors and Simple Motion Forces Circular and Projectile Motion Statics Center of Gravity Work and Energy Power and Momentum Stress and Strain Elasticity and Density Hydrostatic Pressure Fluids in Motion Electricity and Magnetism Plates Capacitors Voltage Batteries Resistors Magnetic Fields Waves and Periodic Motion Sound Waves Doppler Effect Simple Harmonic Motion Optics Mirrors 205 pages Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with

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