

REDOX REACTION

Solved Examples

Example - 6 :



Base Acid

Sol. Valency factor of base = 1

Here, two molecule of NaOH replaced 2H^+ ion from the H_2SO_4 . Therefore, each molecule of NaOH replaced only one H^+ ion of acid, so v.f. = 1.

○ v. f. for acid is the number of OH^- replaced from the base by each molecule of acid

Solved Examples

Example - 7 :



Base Acid

Sol. Valency factor of acid = 1

Here, one of molecule of H_2SO_4 replaced one OH^- from NaOH. Therefore, valency factor for H_2SO_4 is one

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

○ Salts :

(a) In non-reacting condition

○ Valency factor = Total number of positive charge or negative charge present in the compound.

Solved Examples

Example - 8 :

		Na_2CO_3	$\text{Fe}_2(\text{SO}_4)_3$	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
Sol.	v.f.	2	$2 \times 3 = 6$	2
	Eq. wt.	$\frac{M}{2}$	$\frac{M}{6}$	$\frac{M}{2}$

(b) In reacting condition

Solved Examples

Example - 9 :



Base Acid

Sol. It is an acid base reaction, therefore valency factor for Na_2CO_3 is one while in non-reacting condition, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

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Moore

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Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed Robert K. Wismer, 1982

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Study Guide with Solutions for Organic Chemistry, Third Edition Ralph J. Fessenden, Joan S. Fessenden, 1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques

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