

ACIDS, BASES AND SALTS

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|--|--|
| Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions | What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water. |
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|--|--|
| Another term to refer to an acid An acid is called a <u>proton donor</u> | Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction. |
|--|--|

| | | | |
|--|---|---|---|
| Substance to which an acid provides protons BASE | A base is then called PROTON ACCEPTOR | Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids | Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$ |
|--|---|---|---|

| Common laboratory acids | | These three common laboratory acids are also called | Why these acids are also called mineral acids? |
|-------------------------|-----------|---|---|
| Hydrochloric acid | HCl | MINERAL ACIDS | They are derived from mineral salts i.e chlorides for HCl, sulphates for H_2SO_4 and nitrates for HNO_3 |
| Sulphuric acid | H_2SO_4 | | |
| Nitric acid | HNO_3 | | |

| Other mineral acids known | | Mineral salts from which the acid is derived | Organic acids known | Naturally occurring acids |
|---------------------------|-----------|--|------------------------------|---|
| Sulphurous acid | H_2SO_3 | Derived from sulphites | Ethanoic acid (CH_3COOH) | CITRIC ACID from lemons |
| Carbonic acid | H_2CO_3 | Derived from carbonates | | TARTARIC ACID from grapes |
| Phosphoric acid | H_3PO_4 | Derived from phosphates | Methanoic acid ($HCOOH$) | ACETIC ACID from vinegar |
| Nitrous acid | HNO_2 | Derived from nitrites | | LACTIC ACID from sour milk |
| | | | | Hydrochloric acid from digestive juices |

| | | |
|---|---|---|
| Whenever an acid is dissolved in water, it produces HYDROGEN IONS | Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID | Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> . |
|---|---|---|

Chapter 1acids And Bases Study Guide

Joan I. Senyk, Darrell D. Ebbing



Chapter 1acids And Bases Study Guide:

Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry Iris Stovall, Steven S. Zumdahl, 1996

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1. Understanding the eBook Chapter 1acids And Bases Study Guide
 - The Rise of Digital Reading Chapter 1acids And Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1acids And Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1acids And Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1acids And Bases Study Guide
 - Personalized Recommendations
 - Chapter 1acids And Bases Study Guide User Reviews and Ratings
 - Chapter 1acids And Bases Study Guide and Bestseller Lists
5. Accessing Chapter 1acids And Bases Study Guide Free and Paid eBooks

- Chapter 1acids And Bases Study Guide Public Domain eBooks
- Chapter 1acids And Bases Study Guide eBook Subscription Services
- Chapter 1acids And Bases Study Guide Budget-Friendly Options
- 6. Navigating Chapter 1acids And Bases Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 1acids And Bases Study Guide Compatibility with Devices
 - Chapter 1acids And Bases Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1acids And Bases Study Guide
 - Highlighting and Note-Taking Chapter 1acids And Bases Study Guide
 - Interactive Elements Chapter 1acids And Bases Study Guide
- 8. Staying Engaged with Chapter 1acids And Bases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1acids And Bases Study Guide
- 9. Balancing eBooks and Physical Books Chapter 1acids And Bases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1acids And Bases Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1acids And Bases Study Guide
 - Setting Reading Goals Chapter 1acids And Bases Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1acids And Bases Study Guide
 - Fact-Checking eBook Content of Chapter 1acids And Bases Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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