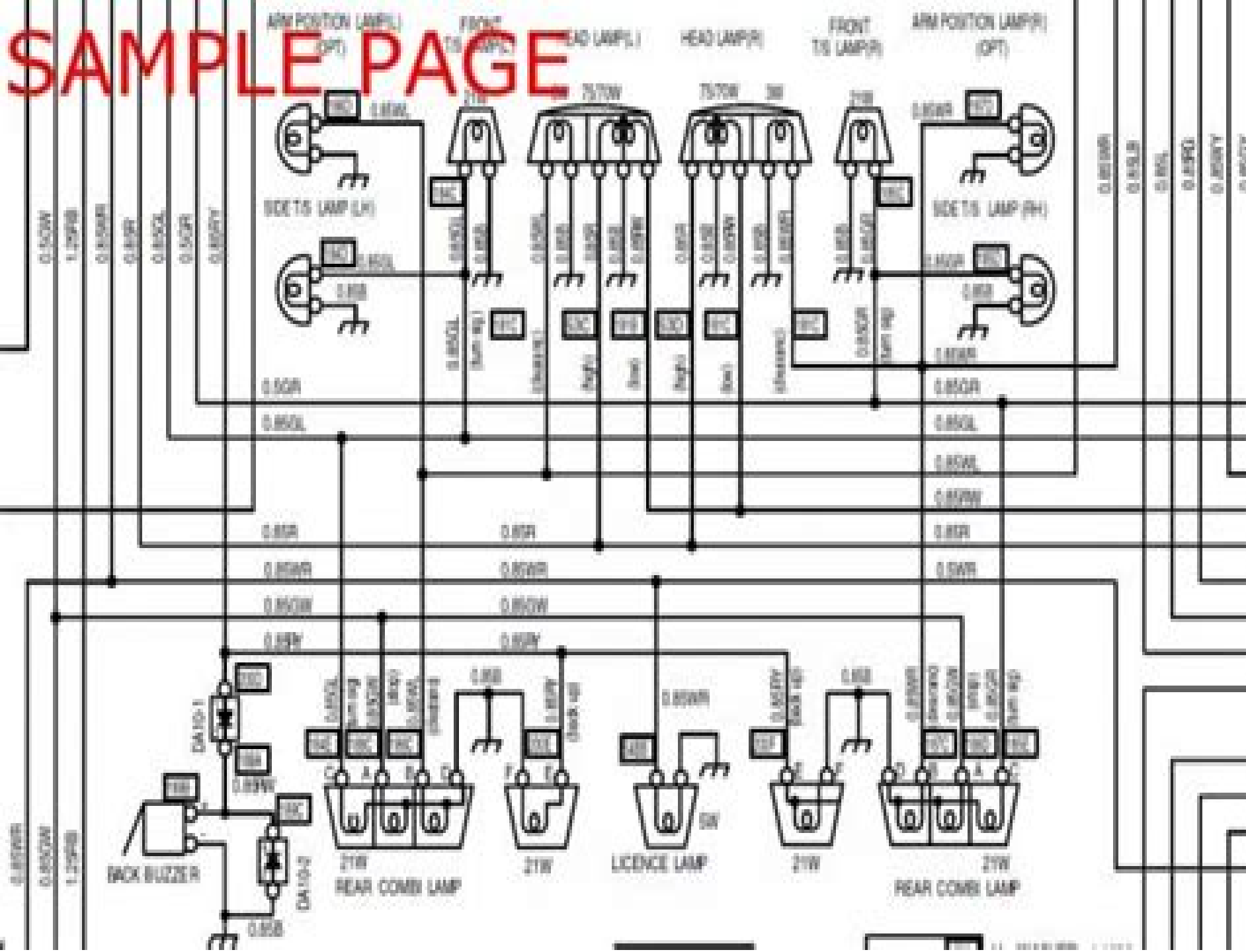


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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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$\text{hcl aq naoh aq nacl aq h}_2\text{o l hcl a q naoh a q nacl a q h}_2\text{o l}$ one mole of hcl hcl would be fully neutralized by one mole of naoh naoh

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{h}_2\text{c}_2\text{o}_4 \cdot 2\text{naoh aq na}_2\text{c}_2\text{o}_4 \cdot 4\text{h}_2\text{o}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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 naoh aq c6h4 cooh cook aq c6h4 coona cook aq h2o l the point of equivalence is the point of titration when the number of
 moles of naoh na added is equal to the number of moles of khp nb in the

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