

# WACE CHEMISTRY NOTES 2016

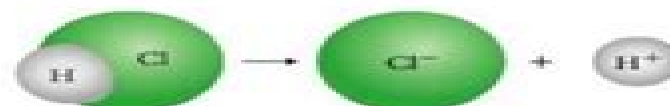
## Intro

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## Acid

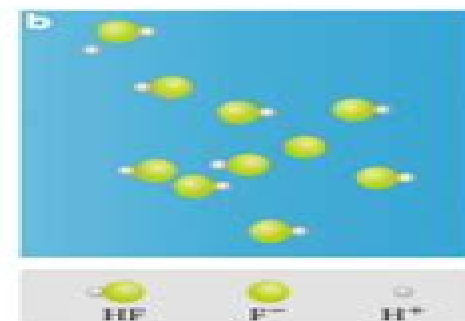
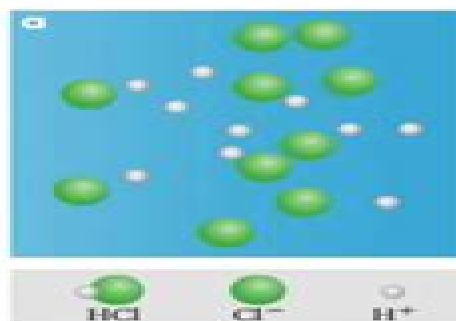
- Produces hydrogen ions
- Or specifically hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous solution
- Reaction with acid + water  $\Rightarrow$  ions are formed  $\Rightarrow$  ionisation reaction
- Strong acids  $\Rightarrow$  all acid molecules ionise  $\Rightarrow$  hydrochloric acid (HCl), Nitric Acid ( $\text{HNO}_3$ ), sulfuric acid ( $\text{H}_2\text{SO}_4$ )

Figure C2.2 ►  
Strong acid ionisation:  
 $\text{HCl} \rightarrow \text{Cl}^- + \text{H}^+$



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Figure C2.3 ►  
Comparison of how  
strong acid (HCl) and  
weak acid (HF) molecules  
ionise in water. a) All of the  
HCl molecules are ionised.  
b) Only some HF molecules  
are ionised.



## Base

- Produces hydroxide ion in solution ie. Soluble salts with oxide ( $\text{O}^{2-}$ ) or hydroxide ( $\text{OH}^-$ ) ion
- Soluble base  $\Rightarrow$  alkali
- Strong bases always dissociate in water to form separate ions  $\Rightarrow$  dissociation reaction

# Dot Point Wace Chemistry

**Mike Lucarelli, David V. Proctor**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the authors' names.

## **Dot Point Wace Chemistry:**

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