

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + \text{Al}(\text{OH})_3(\text{s})$	NONE
2) $\text{NaCl}(\text{aq}) + \text{NH}_4\text{NO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$	NONE
3) $\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{AgI}(\text{s})$	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
4) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + (\text{Cu}_3(\text{PO}_4)_2(\text{s}))$	$3\text{Cu}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Cu}_3(\text{PO}_4)_2(\text{s})$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

Double Replacement Reactions And Solubility Post Lab Questions

Tracy Lynn Haroff



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Table of Contents Double Replacement Reactions And Solubility Post Lab Questions

1. Understanding the eBook Double Replacement Reactions And Solubility Post Lab Questions
 - The Rise of Digital Reading Double Replacement Reactions And Solubility Post Lab Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Double Replacement Reactions And Solubility Post Lab Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Double Replacement Reactions And Solubility Post Lab Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Double Replacement Reactions And Solubility Post Lab Questions
 - Personalized Recommendations
 - Double Replacement Reactions And Solubility Post Lab Questions User Reviews and Ratings
 - Double Replacement Reactions And Solubility Post Lab Questions and Bestseller Lists
5. Accessing Double Replacement Reactions And Solubility Post Lab Questions Free and Paid eBooks
 - Double Replacement Reactions And Solubility Post Lab Questions Public Domain eBooks
 - Double Replacement Reactions And Solubility Post Lab Questions eBook Subscription Services
 - Double Replacement Reactions And Solubility Post Lab Questions Budget-Friendly Options

6. Navigating Double Replacement Reactions And Solubility Post Lab Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Double Replacement Reactions And Solubility Post Lab Questions Compatibility with Devices
 - Double Replacement Reactions And Solubility Post Lab Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Double Replacement Reactions And Solubility Post Lab Questions
 - Highlighting and Note-Taking Double Replacement Reactions And Solubility Post Lab Questions
 - Interactive Elements Double Replacement Reactions And Solubility Post Lab Questions
8. Staying Engaged with Double Replacement Reactions And Solubility Post Lab Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Double Replacement Reactions And Solubility Post Lab Questions
9. Balancing eBooks and Physical Books Double Replacement Reactions And Solubility Post Lab Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Double Replacement Reactions And Solubility Post Lab Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Double Replacement Reactions And Solubility Post Lab Questions
 - Setting Reading Goals Double Replacement Reactions And Solubility Post Lab Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Double Replacement Reactions And Solubility Post Lab Questions
 - Fact-Checking eBook Content of Double Replacement Reactions And Solubility Post Lab Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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