

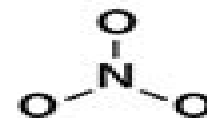
How to Draw Lewis Structure

For Nitrate ion (NO_3^-)

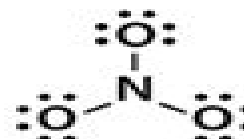
- 1 Determine the number of valence electrons



- 2 Write the skeleton structure



- 3 Fill in the valence electrons



- 4 Transfer electron lone pairs

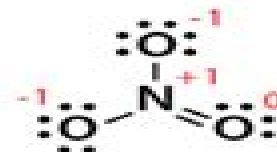


- 5 Determine the formal charge

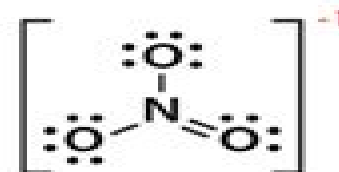
O with single bonds: $6 - 6 - 2/2 = -1$

O with a double bond: $6 - 4 - 4/2 = 0$

N: $5 - 0 - 8/2 = +1$



- 6 Draw a bracket and put the charge



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Table of Contents Drawing Electron Dot Diagrams

1. Understanding the eBook Drawing Electron Dot Diagrams
 - The Rise of Digital Reading Drawing Electron Dot Diagrams
 - Advantages of eBooks Over Traditional Books
2. Identifying Drawing Electron Dot Diagrams
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Drawing Electron Dot Diagrams
 - User-Friendly Interface
4. Exploring eBook Recommendations from Drawing Electron Dot Diagrams
 - Personalized Recommendations
 - Drawing Electron Dot Diagrams User Reviews and Ratings
 - Drawing Electron Dot Diagrams and Bestseller Lists
5. Accessing Drawing Electron Dot Diagrams Free and Paid eBooks
 - Drawing Electron Dot Diagrams Public Domain eBooks
 - Drawing Electron Dot Diagrams eBook Subscription Services
 - Drawing Electron Dot Diagrams Budget-Friendly Options

6. Navigating Drawing Electron Dot Diagrams eBook Formats
 - ePub, PDF, MOBI, and More
 - Drawing Electron Dot Diagrams Compatibility with Devices
 - Drawing Electron Dot Diagrams Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Drawing Electron Dot Diagrams
 - Highlighting and Note-Taking Drawing Electron Dot Diagrams
 - Interactive Elements Drawing Electron Dot Diagrams
8. Staying Engaged with Drawing Electron Dot Diagrams
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Drawing Electron Dot Diagrams
9. Balancing eBooks and Physical Books Drawing Electron Dot Diagrams
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Drawing Electron Dot Diagrams
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Drawing Electron Dot Diagrams
 - Setting Reading Goals Drawing Electron Dot Diagrams
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Drawing Electron Dot Diagrams
 - Fact-Checking eBook Content of Drawing Electron Dot Diagrams
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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