

PROBLEMS WITH CALCULATIONS

1. Nitrogen gas and hydrogen gas combine according to the balanced equation below:



The diagram below shows an initial mixture of nitrogen gas (blue molecules) and hydrogen gas (two small grey spheres each). Before the synthesis of ammonia, explain what happens.



Explanation:

There is only 1 molecule of N_2 gas, consisting of 2 molecules of N . And 3 molecules of H_2 gas, each of 2 molecules of H . So, that adds up to 2 molecules of N and 6 molecules of H . In the final mixture, there are 2 molecules of NH_3 gas, each of 1 molecule of N and 3 molecules of H .

2. Use the same equation as question 1 to answer the following questions.

- a. If 28.0 g of N_2 and 4.0 g of H_2 are reacted, what mass of NH_3 can be produced?

$$\begin{aligned} 28.0 \text{ g } \text{N}_2 \times \frac{1 \text{ mol } \text{N}_2}{28.0 \text{ g } \text{N}_2} &= \frac{1 \text{ mol } \text{N}_2}{28.0 \text{ g } \text{N}_2} \times \frac{2 \text{ mol } \text{NH}_3}{1 \text{ mol } \text{N}_2} = \frac{2 \text{ mol } \text{NH}_3}{28.0 \text{ g } \text{N}_2} \\ 4.0 \text{ g } \text{H}_2 \times \frac{1 \text{ mol } \text{H}_2}{2.0 \text{ g } \text{H}_2} &= \frac{2 \text{ mol } \text{H}_2}{2.0 \text{ g } \text{H}_2} \times \frac{2 \text{ mol } \text{NH}_3}{3 \text{ mol } \text{H}_2} = \frac{4 \text{ mol } \text{NH}_3}{3.0 \text{ g } \text{H}_2} \\ &= 1.33 \text{ mol } \text{NH}_3 \end{aligned}$$

- b. What is the limiting reagent? Explain your work.

Answer: H_2 is the limiting reagent.

Answer: H_2

Answer: H_2

- c. What is the percent yield if 34.0 grams of ammonia is produced in this reaction?

$$\frac{34.0 \text{ g } \text{NH}_3}{100 \text{ g}}$$

Answer: 34.0%

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