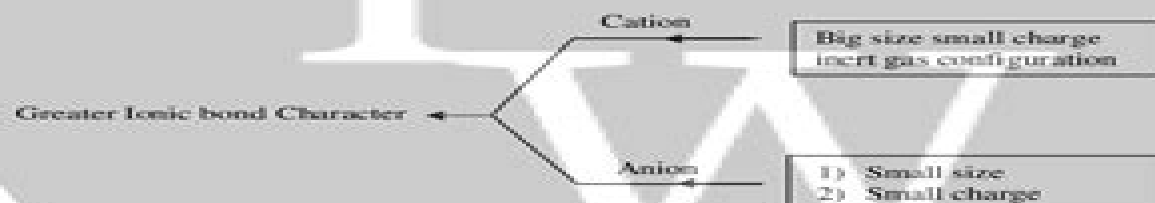


CHAPTER

04

CHEMICAL BONDING & MOLECULAR STRUCTURE

- An atom (or) group of atoms having independent existence in nature is defined as molecule.
- Force of attraction that holds two atoms (or) oppositely charged ions in molecule is chemical bond.
- An atom must have eight electrons in its outermost energy level for its stability is called as octet rule.
- The number of electrons an atom loses (or) gains in an ionic bond formation is called electrovalency.
- According to Fajan's Rules
 - The no. of electrons an atom of the element loses (or) gains is known as electrovalency of that element.
 - Favourable conditions for cation formation.
 - 1) Low I.P
 - 2) Low charge on ion
 - 3) Big atomic size
 - 4) Inert gas configuration
 - Favourable conditions for anion formation.
 - 1) High E.N & E.A
 - 2) Small atomic size
 - 3) Low charge on ion



- LiCl is more covalent in nature and so dissolved in non-polar solvents like alcohol (or) ether.
- Ions which have 18 electrons $[ns^2 np^6 nd^{10}]$ in their outermost main energy level are stable and it is called as pseudo inert gas configuration (or) nickel group configuration.
- The energy liberated when an ionic crystal is formed from an assembly of isolated gaseous ions is called lattice energy.
- Lattice energy

$$(\cup) = \left[-N_0 \frac{AZ^+ Z^- e^2}{r} + N_0 \frac{Be^2}{r^2} \right]$$
 - $$\cup = \frac{AZ^+ Z^- e^2 N_0}{\epsilon_0} \left[1 - \frac{1}{n} \right]$$

(Born – Lande equation)

Coordination number of Na^+ and Cl^- in NaCl crystal is (6:6)

Coordination number of Cs^+ and Cl^- in NaCl crystal is (8:8)
 - The total amount of energy change in a reaction remains the same whether the reaction take place in one (or) several stages is known as Hess law of constant heat summation.

Bonding Chemical Formulas

N. D. Epiotis



Bonding Chemical Formulas:

Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science *Fundamentals of Environmental Chemistry, Third Edition* Stanley E. Manahan, 2011-03-05 Written by an expert using the same approach that made the previous two editions so successful *Fundamentals of Environmental Chemistry Third Edition* expands the scope of book to include the strongly emerging areas broadly described as sustainability science and technology including green chemistry and industrial ecology The new edition includes Increased emphasis on the applied aspects of environmental chemistry Hot topics such as global warming and biomass energy Integration of green chemistry and sustainability concepts throughout the text More and updated questions and answers including some that require Internet research Lecturers Pack on CD ROM with solutions manual PowerPoint presentations and chapter figures available upon qualifying course adoptions The book provides a basic course in chemical science including the fundamentals of organic chemistry and biochemistry The author uses real life examples from environmental chemistry green chemistry and related areas while maintaining brevity and simplicity in his explanation of concepts Building on this foundation the book covers environmental chemistry broadly defined to include sustainability aspects green chemistry industrial ecology and related areas These chapters are organized around the five environmental spheres the hydrosphere atmosphere geosphere biosphere and the anthrosphere The last two chapters discuss analytical chemistry and its relevance to environmental chemistry Manahan's clear concise and readable style makes the information accessible regardless of the readers level of chemistry knowledge He demystifies the material for those who need the basics of chemical science for their trade profession or study curriculum as well as for readers who want to have an understanding of the fundamentals of sustainable chemistry in its crucial role in maintaining a livable planet

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Science of Air Frank R. Spellman, 2016-04-19 Hailed on first publication as a masterful review of the topic The Science of Air Concepts and Applications quickly became a standard resource in the field Clearly written and user friendly the second edition continues to provide the scientific underpinnings of the essence of air Major expansions include Air math and physics Air flow parameters **Deciphering the Chemical Code** N. D. Epiotis, 1996 This groundbreaking work the culmination of more than 10 years of research presents a breakthrough theory of chemical bonding across the periodic table Professor Epiotis an internationally known and respected member of the theoretical community challenges the conventional chemical concepts that underlie popular theories of chemical bonding Building on his insight that electron electron repulsion is the single crucial variable that differentiates one chemical system from another the author formulates explains and applies a new approach based on nonorthogonal valence bond methodology that amounts to nothing less than a revolutionary unified theory of chemical bonding across the periodic table This work represents the first post Pauling theory of chemical bonding New theory means new formulae and this work is about new chemical formulae that lead to the self consistent rationalization of existing facts and even more important the design of new chemistry Pharmaceutical and Chemical Problems and

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