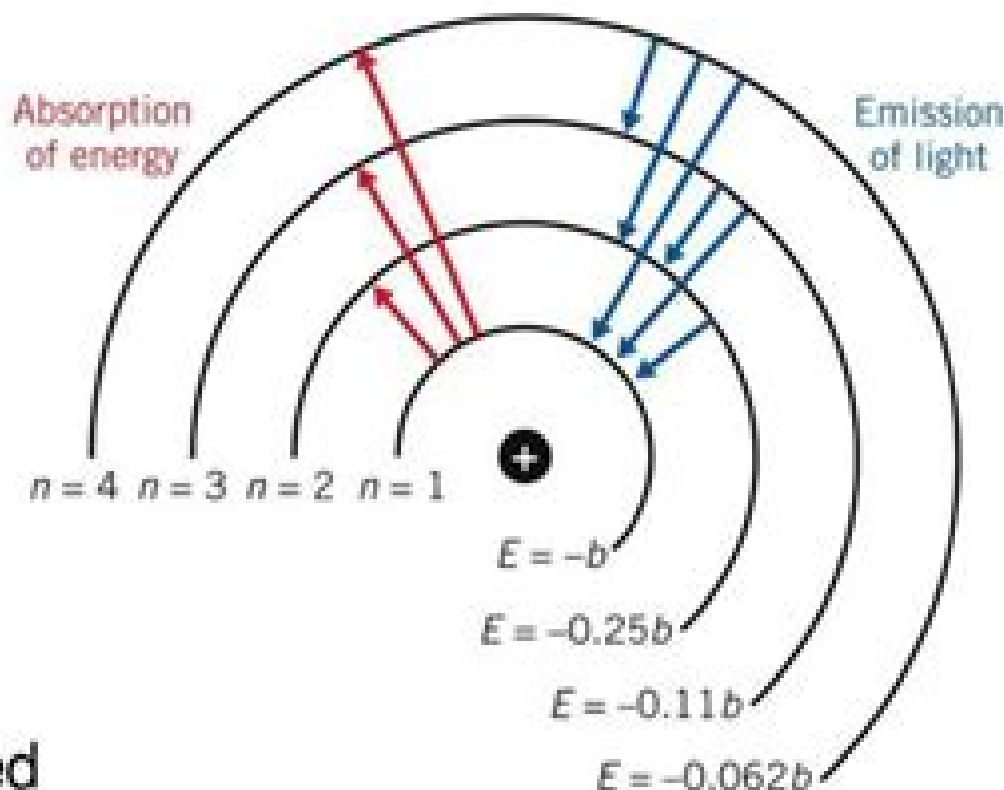


Bohr Model

Energy Level Diagram for a Hydrogen Atom

- Absorption of photon
 - Electron raised to higher energy level
- Emission of photon
 - Electron falls to lower energy level



Energy levels are quantized

- Every time an electron drops from one energy level to a lower energy level
- Same frequency photon is emitted
- Yields line spectra

$$E = -\frac{b}{n^2} = -\frac{R_H hc}{n^2}$$

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Helge Kragh



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