

Practice Problems

10.0001: stoichiometry: stoichiometry calculations

SECTION 12.2: CHEMICAL CALCULATIONS

1. A compound is found to consist of 43.64% oxygen, 56.36% carbon, and 100.00% hydrogen.

$$\text{Empirical formula} = \text{CH}_2\text{O}$$

2. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

$$\text{Empirical formula} = \text{CH}_2\text{O}$$

3. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

$$\text{Empirical formula} = \text{CH}_2\text{O}$$

4. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

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5. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

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6. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

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7. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

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8. A compound is found to consist of 40.00% carbon, 66.67% oxygen, and 33.33% hydrogen.

$$\text{Empirical formula} = \text{CH}_2\text{O}$$

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VM Jensen



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