Name

Bond energy is defined as the amount of energy required to break a bond. These values are positive, indicating that bond breaking is endothermic. Bond energies are reported in kilojoules per mole (kJ/mol). The energy for breaking a hydrogen-hydrogen bond is 436 kJ/mol so when a hydrogen-hydrogen bond is formed the process releases 436 kJ/mol...

In a chemical reaction several bonds are broken and formed. For example in the reaction below a hydrogen-hydrogen bond is broken and a fluorine-fluorine bond is broken. Two hydrogen-fluorine bonds are formed. The overall energy change for this process is calculated below.

H-HBland energy: 436 kJ/mol

F-FBond Energy: 155 kJ/mol

example
$$H_2(g) + F_2(g) \rightarrow 2HF$$

$$\Delta H_m = [\text{energy used for breaking bonds}] - [\text{energy formed in making bonds}]$$

$$[436 \text{ kJ/mol} + 155 \text{ kJ/mol}] - [2(567 \text{ kJ/mol})]$$

$$= -543 \text{ kJ/mol}$$

Average Bond Energies (kJ/mol)

H-H	4.36 kJ/mod	C-H	413 kl/mol	C=C	614 kJ//mol
H-C1	431 kJ/mol	C-C	348 kil/mol	CHC	839 kJVmol
H-F	567 kJ/mol	C-N	293 kJ/mol	C=O	799 kJ//mol
N-H	3/91 kJ/molt	C-O	3.58 kJVmod	O=O	495 kJ//mol
N-O	201 kJ/mol	C-F	485 kJ/mod	CHO	1072 kJ/mod
O-H	463 k3/mol	C-CI	328 kJ/mol	CHN	615 k.Wmell
O-O	146 kJ/molt	C-S	2.59 kJ//mol	N-N	418 kJVmol
F-F	155 kJ/mol	CI-CI	242 kJ/mol	NewNo	941 kJ/mol
				CHN	891 kJ/mol

Estimate the enthalpy change (ΔH_{res}) of the following reactions using the bond energies above.

1.
$$H-H+CI-CI \rightarrow H-CI+H-CI$$
4. $H-\frac{CI}{C}-CI+O \rightarrow \frac{O}{C}-CI+H-CI$
2. $H-C=C + H-F-F \rightarrow H-\frac{H}{C}-\frac{H}{C}-H$
5. $H-\frac{C}{C}-H+O=O \rightarrow O=C=O+H-O-H$
3. $H-\frac{H}{C}-O-H+H-\frac{H}{C}-O-H \rightarrow H-\frac{H}{C}-O-H+H-O-H$

Draw Lewis structures for reactants and products. Estimate the enthalpy change (ΔH_{re}) for the reactions using bond energies.

9.
$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

10.
$$H_2(g) + C_2H_4(g) \rightarrow C_2H_6(g)$$

11.
$$2C_2H_6(g) + 7O_2(g) \rightarrow 4CO_2(g) + 6H_2O(g)$$

13. HCN
$$(g)$$
 + 2H₂ (g) \rightarrow CH₃NH₂ (g)

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Ying-Ying Zheng

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