

## Chapter 6: Chemical Bonding Guided Notes

### I. Why do Atoms Combine?

#### A. Chemical Formula

- The chemical formula shows:
  - The \_\_\_\_\_ in the compound.
  - The ratio of their \_\_\_\_\_

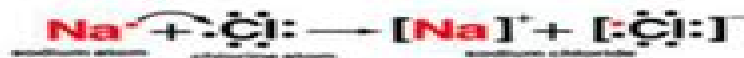


#### B. Chemical Bond

- Strong attractive force between \_\_\_\_\_ or ions in a molecule or compound.
- Formed by:
  - \_\_\_\_\_ electrons, (e-) by losing or gaining electrons.
  - \_\_\_\_\_ electrons (e-)

#### C. Stability

- Octet Rule
  - Most atoms form bonds in order to have \_\_\_\_\_ valence electrons (e-).
  - \_\_\_\_\_ outer energy level.
  - Want to be like the Noble Gases (group 18).
- Stability is the driving force behind bond formation. All atoms bond to become stable.
- Transferring e-



- Sharing e-



### II. Kinds of Chemical Bonds

#### A. Ionic Bond

- Attraction between 2 \_\_\_\_\_ charged ions.
- Ions are \_\_\_\_\_ atoms.
- A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom loses an electron.  
Formed by \_\_\_\_\_.
- A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom gains an electron.  
Formed by \_\_\_\_\_.
- Formed by \_\_\_\_\_ electrons (e-) from a \_\_\_\_\_ to a \_\_\_\_\_.
- Ions form a 3-D crystal lattice.



# Bonding Guided Notes Answers

**Johnny P. Bradley**



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## **Table of Contents Bonding Guided Notes Answers**

1. Understanding the eBook Bonding Guided Notes Answers
  - The Rise of Digital Reading Bonding Guided Notes Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Bonding Guided Notes Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bonding Guided Notes Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Bonding Guided Notes Answers
  - Personalized Recommendations
  - Bonding Guided Notes Answers User Reviews and Ratings
  - Bonding Guided Notes Answers and Bestseller Lists
5. Accessing Bonding Guided Notes Answers Free and Paid eBooks

- Bonding Guided Notes Answers Public Domain eBooks
- Bonding Guided Notes Answers eBook Subscription Services
- Bonding Guided Notes Answers Budget-Friendly Options
- 6. Navigating Bonding Guided Notes Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Bonding Guided Notes Answers Compatibility with Devices
  - Bonding Guided Notes Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Bonding Guided Notes Answers
  - Highlighting and Note-Taking Bonding Guided Notes Answers
  - Interactive Elements Bonding Guided Notes Answers
- 8. Staying Engaged with Bonding Guided Notes Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bonding Guided Notes Answers
- 9. Balancing eBooks and Physical Books Bonding Guided Notes Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bonding Guided Notes Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Bonding Guided Notes Answers
  - Setting Reading Goals Bonding Guided Notes Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bonding Guided Notes Answers
  - Fact-Checking eBook Content of Bonding Guided Notes Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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