

## Addition Compounds

Double Salt.

Complex Salt.

\* Carnallite :-



\* Mohr's Salt :-



\* Potash Alum :-



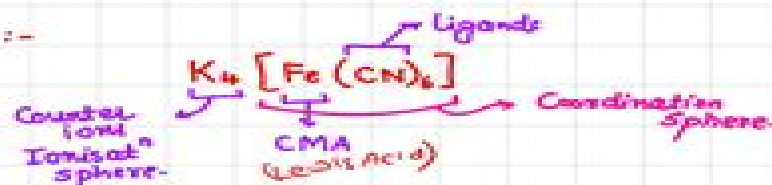
\* d-block elements.

↳ \* Small Size,

\* high  $\phi$  Nature.

\* Vacant Orbitals

Note :-



\* **Denticity** :- No. of  $e^-$  pairs which can be donated by a ligand to metal ion.

\* In any ligand if total no. of atoms are 3 or less than 3  $\rightarrow$  then denticity is always "1".

Stable  $\rightarrow$  free

## CLASSIFICATION OF LIGANDS :-

A) Anionic

$\text{SO}_4^{2-}$  = Sulfate

$\text{CO}_3^{2-}$  = Carbonate

$\text{Cl}^-$  = Chloride

$\text{OH}^-$  = Hydroxide

$\text{F}^-$  = Fluoride

$\text{NH}_2^-$  = Amide

$\text{NH}^{2-}$  = Imide

$\text{N}_3^-$  = Azide

$\text{N}^{3-}$  = Nitride

$\text{O}^{2-}$  = Oxide

$\text{O}_2^-$  = Superoxide

$\text{O}_2^{2-}$  = Peroxide

$\text{CN}^-$  = Cyanide

$\text{NC}^-$  = Isocyanide

$\text{NO}_2^-$  = Nitrite - N

$\text{ONO}^-$  = Nitrite - O

$\text{OCN}^-$  = Cyanate - O

$\text{NCO}^-$  = Cyanate - N

$\text{SCN}^-$  = Thiocyanate - S

$\text{NCS}^-$  = Thiocyanate - N

$\text{C}_2\text{O}_4^{2-}$  = Oxalate (2)

Gly = Glycinate (2)

dimethylgly = dimethylglycinate (2)

glyoximate (2)

$\text{EDTA}^{4-}$  = Ethylene diamine tetraacetate. (6)

B) Neutral :-

$\text{H}_2\text{O}$  :- Aqua

$\text{NH}_3$  :- Amine

$\text{CO}$  :- Carbonyl

$\text{NO}$  :- Nitrosyl

$\text{PH}_3$  :- Phosphine

$\text{PPh}_3$  :- Triphenyl Phosphine

$\text{Py}$  :- Pyridine.

en :- ethylenediamine.

pn :- propylene diamine

tn :- trimethylene diamine

dien :- diethylene triamine

trien :- triethylene tetraamine.

C) Cationic :-

$\text{NO}^+$  = Nitrosonium

$\text{NH}_2^+ - \text{NH}_3^+$  :- Hydrazonium.

Jugaad :-

$\rightarrow \text{Pt/Pd}$

always  $d^8$

$\text{CN} = 4$

O.S. = +2

$\text{CN} = 6$

O.S. = +4.

$\rightarrow$  When Fe & NO both +ve.

$\text{NO} = +1$

$\rightarrow \text{Sc Ti V Cr Mn Fe Co Ni Cu Zn} \}^{+2}$

3d = 1 2 3 4 5 6 7 8 9 10

Similar for below 4d & 5d !

$\rightarrow d^1/d^2/d^3 \xrightarrow[6 \text{ ligands}]{\text{SL/OL}} d^2sp^3$   
 $\text{UPE}^- = 1 \ 2 \ 3$   $\text{CN} = 6$  always Ioc LSC Paramag.

$\rightarrow d^2/d^3/d^4 \xrightarrow[6 \text{ ligands}]{\text{SL/OL}} sp^3d^2$   
 $\text{UPE}^- = 2 \ 1 \ 0$   $\text{CN} = 6$  always ooc HSC Para Para dia.

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1. Understanding the eBook Class 1coordination Compounds Notes From
  - The Rise of Digital Reading Class 1coordination Compounds Notes From
  - Advantages of eBooks Over Traditional Books
2. Identifying Class 1coordination Compounds Notes From
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Class 1coordination Compounds Notes From
  - User-Friendly Interface
4. Exploring eBook Recommendations from Class 1coordination Compounds Notes From
  - Personalized Recommendations

- Class 1coordination Compounds Notes From User Reviews and Ratings
- Class 1coordination Compounds Notes From and Bestseller Lists
- 5. Accessing Class 1coordination Compounds Notes From Free and Paid eBooks
  - Class 1coordination Compounds Notes From Public Domain eBooks
  - Class 1coordination Compounds Notes From eBook Subscription Services
  - Class 1coordination Compounds Notes From Budget-Friendly Options
- 6. Navigating Class 1coordination Compounds Notes From eBook Formats
  - ePub, PDF, MOBI, and More
  - Class 1coordination Compounds Notes From Compatibility with Devices
  - Class 1coordination Compounds Notes From Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Class 1coordination Compounds Notes From
  - Highlighting and Note-Taking Class 1coordination Compounds Notes From
  - Interactive Elements Class 1coordination Compounds Notes From
- 8. Staying Engaged with Class 1coordination Compounds Notes From
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Class 1coordination Compounds Notes From
- 9. Balancing eBooks and Physical Books Class 1coordination Compounds Notes From
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Class 1coordination Compounds Notes From
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Class 1coordination Compounds Notes From
  - Setting Reading Goals Class 1coordination Compounds Notes From
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Class 1coordination Compounds Notes From
  - Fact-Checking eBook Content of Class 1coordination Compounds Notes From

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
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