Double Replacement Reactions and Solubility

Post-Lab Questions: Review you Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1)	NI- OO	MANO	→ 6NaNO3(09)+M(04)3(M9)	None
2)		NUNO	> HONOS(04) + NHACI (04)	None
35			100 ton (00) + 40 T(0)	Agi (ag) + I (ag) + AgI(s)
49	Nat por + NasPOrms +	C March 1997 Comment	+ 6NaNoz(ag)+(mz(189)z(5)	36+21(09) + 2 PO((44) + 605(PO4) 2(5)

NasPOvis	+ Cu(NO ₃) _{Nong} → l ₂ N c	1 NO3(04)+ (W2(104)2(5))	T. Cally Charles
Use your obse	ervations table to answer the	following. ALL ions must include	a symbol AND a charge or it is wrong!
		is that produced the fewest (<3) pre	
		d as SOLUBLE except for those w	
		that produced the most (>5) precip	
		d as INSOLUBLE except for those	
7. Although y to be solul	you did not test the bromide ble or insoluble? Hint: Lo	ion (Br '), would you expect most be cate bromine on the periodic table.	romide salts (AlBr., BaBr., CaBr., etc)
	our answer in #7		
		an exception to the rule in #7?	
10-14 Summa	rize the general solubility ru	les of ionic compounds in the table	below. Two have been done for you.
	lon	Generally soluble or insoluble?	Exceptions
	9) carbonates		The second secon
	 chlorides, bromides, iodides 		
	11) hydroxides		
	nitrates	Soluble	None

12) phosphates

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Vasyl Andrushko, Natalia Andrushko

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Table of Contents Double Replacement Reaction And Solubility Lab Key

- 1. Understanding the eBook Double Replacement Reaction And Solubility Lab Key
 - The Rise of Digital Reading Double Replacement Reaction And Solubility Lab Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Double Replacement Reaction And Solubility Lab Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Double Replacement Reaction And Solubility Lab Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Double Replacement Reaction And Solubility Lab Key
 - Personalized Recommendations
 - $\circ\,$ Double Replacement Reaction And Solubility Lab Key User Reviews and Ratings
 - Double Replacement Reaction And Solubility Lab Key and Bestseller Lists
- 5. Accessing Double Replacement Reaction And Solubility Lab Key Free and Paid eBooks
 - Double Replacement Reaction And Solubility Lab Key Public Domain eBooks
 - Double Replacement Reaction And Solubility Lab Key eBook Subscription Services
 - Double Replacement Reaction And Solubility Lab Key Budget-Friendly Options
- 6. Navigating Double Replacement Reaction And Solubility Lab Key eBook Formats

- o ePub, PDF, MOBI, and More
- Double Replacement Reaction And Solubility Lab Key Compatibility with Devices
- Double Replacement Reaction And Solubility Lab Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Double Replacement Reaction And Solubility Lab Key
 - Highlighting and Note-Taking Double Replacement Reaction And Solubility Lab Key
 - Interactive Elements Double Replacement Reaction And Solubility Lab Key
- 8. Staying Engaged with Double Replacement Reaction And Solubility Lab Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Double Replacement Reaction And Solubility Lab Key
- 9. Balancing eBooks and Physical Books Double Replacement Reaction And Solubility Lab Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Double Replacement Reaction And Solubility Lab Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Double Replacement Reaction And Solubility Lab Key
 - Setting Reading Goals Double Replacement Reaction And Solubility Lab Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Double Replacement Reaction And Solubility Lab Key
 - Fact-Checking eBook Content of Double Replacement Reaction And Solubility Lab Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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