

Name: \_\_\_\_\_

## Double Replacement Lab

Reminder – Goggles must be worn at all times in the lab!

### PRE-LAB DISCUSSION:

A PRECIPITATE is an insoluble substance formed by the reaction between chemicals in solution. A list of reactions will be given to you. The reactants are on the left side of the equation, and the products are on the right. The precipitate will be found on the product side. Keep in mind that an equation gives a condensed picture of what happens when two chemicals react.

### PURPOSE:

The purpose of the lab is to observe semi-micro scale precipitation reactions, and to use a solubility table to identify the insoluble product, when there is one.

### PROCEDURE:

Predict the products and balance the following double replacement reactions. Then, identify the precipitates using the table on page 2. **You may want to do your work step on a scratch piece of paper and attach to lab later.**

<u>Double Replacement Reactions</u>					<b>Precipitates (if any)</b>
A.	NaCl	+	AgNO <sub>3</sub>	→	
B.	KI	+	AgNO <sub>3</sub>	→	
C.	NaOH	+	AgNO <sub>3</sub>	→	
D.	Na <sub>3</sub> PO <sub>4</sub>	+	AgNO <sub>3</sub>	→	
E.	NaCl	+	Pb(NO <sub>3</sub> ) <sub>2</sub>	→	
F.	KI	+	Pb(NO <sub>3</sub> ) <sub>2</sub>	→	
G.	NaOH	+	Pb(NO <sub>3</sub> ) <sub>2</sub>	→	
H.	Na <sub>3</sub> PO <sub>4</sub>	+	Pb(NO <sub>3</sub> ) <sub>2</sub>	→	
I.	NaCl	+	CuSO <sub>4</sub>	→	
J.	KI	+	CuSO <sub>4</sub>	→	
K.	NaOH	+	CuSO <sub>4</sub>	→	
L.	Na <sub>3</sub> PO <sub>4</sub>	+	CuSO <sub>4</sub>	→	

# Double Replacement Reactions Lab 9 2 Key

**Scott C. Dulebohn**



## Double Replacement Reactions Lab 9 2 Key:

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