

Science 10 Review Answers
Chemistry Review

1. Protons, neutrons and, electrons are the three particles that make up the atom. Protons and neutrons make up 99.9% of the mass equally, electrons make up very little of the mass. Protons and neutron are found in the nucleus and electrons orbit the nucleus in energy levels. Protons have a positive charge, neutrons have no charges and, electrons have a negative charge.
2. Name the:
 - a. Element with 8 protons - Oxygen
 - b. Element with 10 electrons - Neon
 - c. Element with 3 electron shells (energy levels) and 5 electrons in the valence energy level - Phosphorous
 - d. Ion with a 2+ charge containing 10 electrons - Magnesium
 - e. Ion with a 2- charge containing 18 electrons - Sulfur
 - f. Ion with 38 protons and 36 electrons – strontium ion
3. Multi-valent elements or multi-valent metals.
4. Provide the symbol or name for the following:
 - a. Cu – copper atom
 - b. Cu^{2+} - copper (II) ion
 - c. O – oxygen atom
 - d. O^{2-} - oxide ion
 - e. SO_4^{2-} - sulfate ion
 - f. hydroxide ion - OH^-
 - g. zinc atom – $\text{Zn}_{(s)}$
 - h. sulfur atom – $\text{S}_{(s)}$
 - i. calcium ion – Ca^{2+}
 - j. iodide ion - I^-
5. Ionic bonds require the transfer of electron and the attraction between a positive and negative ion. A covalent bond is where two atoms share electrons. Ionic bonds are stronger.
6. Name the following:
 - a. $\text{NaF}_{(s)}$ –sodium fluoride
 - b. $\text{BaCl}_{2(s)}$ –barium chloride
 - c. $\text{CCl}_{4(l)}$ – carbon tetrachloride
 - d. $\text{CaSO}_{4(s)}$ –calcium sulfate
 - e. $\text{O}_{2(g)}$ -oxygen
 - f. $\text{SO}_{2(g)}$ –sulfur dioxide
 - g. $\text{Cu}(\text{OH})_{2(s)}$ –copper (II) hydroxide
 - h. $\text{Al}(\text{NO}_3)_{3(s)}$ –aluminum nitrate
 - i. $\text{Pb}(\text{OH})_{4(s)}$ –lead (IV) hydroxide
 - j. $\text{H}_{2(g)}$ -hydrogen
 - k. $\text{K}_2\text{S}_{(s)}$ –potassium sulfide
 - l. $\text{S}_{8(s)}$ –sulfur or octasulfur
 - m. $\text{H}_3\text{PO}_{4(aq)}$ – phosphoric acid
 - n. $\text{HBr}_{(aq)}$ – hydrobromic acid
 - o. $\text{HClO}_{2(aq)}$ – chlorous acid
 - p. $\text{NH}_{3(g)}$ - ammonia
 - q. $\text{C}_6\text{H}_{12}\text{O}_{6(s)}$ - glucose
 - r. $\text{H}_2\text{S}_{(g)}$ - hydrogen sulfide
7. Write the formulae for the following :
 - a. copper (I) oxide – $\text{Cu}(\text{OH})_{(s)}$
 - b. sodium iodide – $\text{NaI}_{(s)}$
 - c. zinc nitrate – $\text{Zn}(\text{NO}_3)_{2(s)}$

Bc Science 10 Review Answer Key

G Thomas



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